# Stoichiometry: Mass – Mole Worksheet

Solve the following stoichiometry problems and type in the correct answer with units. (for example: 3.45 mol or 4.62 g) Round the masses from the periodic table to the nearest whole number. Round all answers to 2 decimal places.

1. Based on the following chemical equation, please answer the following questions

Fe (SO<sub>4</sub>)<sub>3</sub> + 6 NaCl → 2 FeCl<sub>3</sub> + 3 Na<sub>2</sub>SO<sub>4</sub>

Calculate the amount of moles of NaCl needed to produce 75 g of FeCl<sub>3</sub>.

Answer: g FeCl3 | Mal FeCl3 | Mal Nal 1 = g FeCl3 | mal FeCl3

b. Calculate the amount of Fe<sub>2</sub>(SO<sub>4</sub>)<sub>3</sub> in grams produce 215 g Na<sub>2</sub>SO<sub>4</sub>.

Answer: \_\_\_\_\_ g Nas Say | mal Nas Say | mal

Calculate the amount in grams of FeCl<sub>3</sub> produced with 95g of Fe(SO<sub>4</sub>)<sub>3</sub>.

Answer: gFe (SA) 3 | Mol Fe (S

Calculate the amount of moles of Fe<sub>2</sub>(SO<sub>4</sub>)<sub>3</sub> needed to react with 123.5 g of NaCl.

Answer: \_\_\_\_\_\_ g Nacl | mol Nacl | mol Fen(soly = g Nacl | mol Nacl

How many moles of CaCO<sub>3</sub> can be produced if we make 12.3 g of Ca(HCO<sub>3</sub>)<sub>2</sub> reacts with CaCO<sub>3</sub>?

 $Ca(HCO_1) + Ca(OH)_2 \rightarrow 2 CaCO_3 + 2 H_1O$ 

Answer: g(a(HCOs)2) | Mol (a(HCOs)2) Mol (a COs = g Ca(HCOs)2 | Mol (a(HCOs)2

# **Mole Mole And Mass Mass Stoichiometry Worksheet**

**KJ Lindholm-Leary** 

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