

Mole to Grams, Grams to Moles Conversions Worksheet*What are the molecular weights of the following compounds? (all masses must be to nearest hundredth)*

- | | |
|-------------|-------------------|
| 1) NaOH | 2) H_3PO_4 |
| 3) H_2O | 4) Mn_2Se_7 |
| 5) $MgCl_2$ | 6) $(NH_4)_2SO_4$ |

There are three definitions (equalities) of mole. They are:1 mole = 6.02×10^{23} particles

1 mole = molar mass (could be atomic mass from periodic table or molecular mass)

1 mole = 22.4 L of a gas at STP (You do not need to worry about this yet)

Each definition can be written as a set of two conversion factors. They are:1 mole = molar mass(g) can be written as $\left(\frac{1 \text{ mole}}{\text{molar mass (g)}} \right)$ OR $\left(\frac{\text{molar mass (g)}}{1 \text{ mole}} \right)$ 1 mole = 6.02×10^{23} particles can be written as $\left(\frac{1 \text{ mole}}{6.02 \times 10^{23}} \right)$ OR $\left(\frac{6.02 \times 10^{23}}{1 \text{ mole}} \right)$ **Solve the following:**

- 1)
- How many moles are in 15 grams of lithium? (molar mass of lithium is 6.94 g/mole)**

$$15 \text{ grams} \times \frac{1 \text{ mole}}{6.94 \text{ grams}} = 2.1614 \text{ moles lithium} = \boxed{2.2 \text{ moles Li}}$$

- 2)
- How many grams are in 2.4 moles of sulfur? (molar mass of sulfur is 32.07 g/ mole)**

$$2.4 \text{ moles} \times \frac{32.07 \text{ grams}}{1 \text{ mole}} = 76.97 \text{ grams sulfur} = \boxed{77 \text{ g Sulfur}}$$

- 3) **How many moles are in 22 grams of argon?**
- 4) **How many grams are in 88.1 moles of magnesium?**
- 5) **How many moles are in 2.3 grams of phosphorus?**

Mole Conversion Worksheet 1 Answers

Jiyuan Zhang



Mole Conversion Worksheet 1 Answers:

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Table of Contents Mole Conversion Worksheet 1 Answers

1. Understanding the eBook Mole Conversion Worksheet 1 Answers
 - The Rise of Digital Reading Mole Conversion Worksheet 1 Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Mole Conversion Worksheet 1 Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Mole Conversion Worksheet 1 Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Mole Conversion Worksheet 1 Answers
 - Personalized Recommendations
 - Mole Conversion Worksheet 1 Answers User Reviews and Ratings
 - Mole Conversion Worksheet 1 Answers and Bestseller Lists
5. Accessing Mole Conversion Worksheet 1 Answers Free and Paid eBooks
 - Mole Conversion Worksheet 1 Answers Public Domain eBooks

- Mole Conversion Worksheet 1 Answers eBook Subscription Services
- Mole Conversion Worksheet 1 Answers Budget-Friendly Options
- 6. Navigating Mole Conversion Worksheet 1 Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Mole Conversion Worksheet 1 Answers Compatibility with Devices
 - Mole Conversion Worksheet 1 Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Mole Conversion Worksheet 1 Answers
 - Highlighting and Note-Taking Mole Conversion Worksheet 1 Answers
 - Interactive Elements Mole Conversion Worksheet 1 Answers
- 8. Staying Engaged with Mole Conversion Worksheet 1 Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Mole Conversion Worksheet 1 Answers
- 9. Balancing eBooks and Physical Books Mole Conversion Worksheet 1 Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Mole Conversion Worksheet 1 Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Mole Conversion Worksheet 1 Answers
 - Setting Reading Goals Mole Conversion Worksheet 1 Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Mole Conversion Worksheet 1 Answers
 - Fact-Checking eBook Content of Mole Conversion Worksheet 1 Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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