

## Unit: Stoichiometry Answer key "Mole-Mole Calculations" - WS #1

Directions: For each of the following problems, show your balanced chemical equation and all of your calculations.

 How many moles of potassium oxide (K<sub>2</sub>O) will be formed when 1.52 moles of potassium reacts with oxygen according to the following reaction:

4 K + O<sub>2</sub> → 2 K<sub>2</sub>O  
1.52 mol K x 
$$\frac{2 \mod K_2 O}{4 \mod K}$$
 = 0.760 mol K<sub>2</sub>O

How many moles of copper are needed to react with sulfur to produce 0.25 moles of copper(I) sulfide according to the following reaction:

$$\begin{array}{cccc} \textbf{2 Cu} & + \textbf{S} & \rightarrow \textbf{Cu}_2\textbf{S} \\ 0.25 & mol & Cu_2Sx & \frac{2 \ mol \ Cu}{1 \ mol \ Cu_2S} = 0.50 \ mol \ Cu \end{array}$$

How many moles of sodium are needed to react with sulfuric acid to produce 3.75 moles of sodium sulfate according to the following (unbalanced) reaction:

2\_ Na + \_\_\_\_\_ H<sub>2</sub>(SO<sub>4</sub>) 
$$\Rightarrow$$
 \_\_\_\_\_ Na<sub>2</sub>(SO<sub>4</sub>) + \_\_\_\_\_ H<sub>2</sub>
3.75 mol Na<sub>2</sub>SO<sub>4</sub>  $\times$   $\frac{2 \text{ mol Na}}{1 \text{ mol Na}_2SO_4} = 7.50 \text{ mol Na}$ 

How many moles of calcium oxide will be produced when 1.6 moles of Iron(III)oxide reacts with calcium phosphate according to the following (unbalanced) reaction:

\_\_\_\_ Fe<sub>2</sub>O<sub>3</sub> + \_\_\_\_ Ca<sub>3</sub>(PO<sub>4</sub>)<sub>2</sub> 
$$\rightarrow$$
 \_\_\_ 2\_\_ Fe(PO<sub>4</sub>) + \_\_\_ 3\_\_ CaO  
1.6 mol Fe<sub>2</sub>O<sub>3</sub> x  $\frac{3 mol CaO}{1 mol Fe2O3 = 4.8 mol Ca}$ 

4. How many moles of oxygen are produced when 5.12 moles of magnesium chlorate decomposes into magnesium chloride and oxygen?

Word equation: magnesium chlorate → magnesium chloride + oxygen

Formula equation: 
$$Mg(ClO_3)_2 \rightarrow MgCl_2 + 3O_2$$
 (must be balanced!)
$$5.12 \mod Mg(ClO_3)2 \times \frac{3 \mod O_2}{1 \mod Mg(ClO_3)2} = 15.4 \mod O_2$$

How many moles of Chromium(III) nitrate are produced when Chromium reacts with 0.85 moles of Lead(IV) nitrate to produce chromium(III) nitrate and lead?

Word equation: chromium + lead (IV) nitrate → chromium (III) nitrate + lead

Formula equation: 
$$4\text{Cr} + 3\text{Pb}(\text{NO}_3)_4 \Rightarrow 4\text{Cr}(\text{NO}_3)_3 + 3\text{Pb}$$
 (must be balanced!)
$$0.85 \ mol \ Pb(NO_3)_4 \times \frac{4 \ mol \ Cr(NO_3)_3}{3 \ mol \ Pb(NO_3)_4} = 1.1 \ mol \ Cr(NO_3)_3$$

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#### **Table of Contents Mole 8 Answer Key**

- 1. Understanding the eBook Mole 8 Answer Key
  - The Rise of Digital Reading Mole 8 Answer Key
  - Advantages of eBooks Over Traditional Books
- 2. Identifying Mole 8 Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Mole 8 Answer Key
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from Mole 8 Answer Key
  - Personalized Recommendations
  - $\circ\,$  Mole 8 Answer Key User Reviews and Ratings
  - $\circ\,$  Mole 8 Answer Key and Bestseller Lists
- 5. Accessing Mole 8 Answer Key Free and Paid eBooks
  - Mole 8 Answer Key Public Domain eBooks

- Mole 8 Answer Key eBook Subscription Services
- Mole 8 Answer Key Budget-Friendly Options
- 6. Navigating Mole 8 Answer Key eBook Formats
  - o ePub, PDF, MOBI, and More
  - Mole 8 Answer Key Compatibility with Devices
  - Mole 8 Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Mole 8 Answer Key
  - Highlighting and Note-Taking Mole 8 Answer Key
  - Interactive Elements Mole 8 Answer Key
- 8. Staying Engaged with Mole 8 Answer Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Mole 8 Answer Key
- 9. Balancing eBooks and Physical Books Mole 8 Answer Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Mole 8 Answer Key
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Mole 8 Answer Key
  - Setting Reading Goals Mole 8 Answer Key
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Mole 8 Answer Key
  - Fact-Checking eBook Content of Mole 8 Answer Key
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks

## 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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