Name\_\_\_\_\_\_\_Date:

#### The Mole Webquest

Use the following website to answer this set of questions:

http://antoine.frostburg.edu/chem/senese/101/moles/faq/why-use-moles.shtml

1) A mole of anything is how many? (give the number):

 $6.02 \times 10^{22}$ 

 Why is it that different amounts of things can still equal one mole? (think about the weight of a dozen elephants vs a dozen eggs)

The number of grams in a mole is different from substance to substance.

3) Why do we want to use the concept of moles?

You often want to know how many molecules you have in a sample of a substance.

With moles, you count the number of molecules in the sample by weighing it.

Once we know the number of moles we can convert to the number of:
 moles or grams and vice versa.

5) How many grams of water are in one mole of water?

1 mole water = 18 g water

6) How many molecules of water are in one mole of water?

 $6.02 \times 10^{23}$ 

Use the following website to answer the next set of questions: https://chemfiesta.wordpress.com/2015/02/06/mole-calculations-3/

Define molar mass.

The mass of a given substance (chemical element or chemical compound is in g) divided by its amount of substance in (mol).

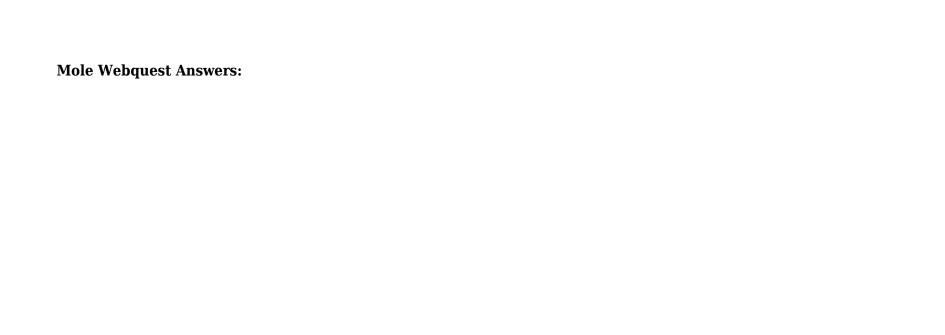
8) What is the mass of one mole of oxygen? 16.0 grams per mole

8) Find the molar mass of NH3NO3. (Show your work)

Ammonium nitrate has two nitrogen atoms (2 x 14.0 g/mol), four hydrogen atoms (4 x 1.0 g/mol), and three oxygen atoms (3 x 16.0 g/mol). This makes the molar mass 80.0 g/mol.

# **Mole Webquest Answers**

JE Gale



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## **Table of Contents Mole Webquest Answers**

- 1. Understanding the eBook Mole Webquest Answers
  - The Rise of Digital Reading Mole Webquest Answers
  - Advantages of eBooks Over Traditional Books
- 2. Identifying Mole Webquest Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Mole Webquest Answers
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from Mole Webquest Answers
  - Personalized Recommendations
  - Mole Webquest Answers User Reviews and Ratings
  - Mole Webguest Answers and Bestseller Lists
- 5. Accessing Mole Webquest Answers Free and Paid eBooks
  - Mole Webquest Answers Public Domain eBooks
  - Mole Webguest Answers eBook Subscription Services
  - Mole Webquest Answers Budget-Friendly Options
- 6. Navigating Mole Webquest Answers eBook Formats

- ∘ ePub, PDF, MOBI, and More
- Mole Webquest Answers Compatibility with Devices
- Mole Webquest Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Mole Webquest Answers
  - Highlighting and Note-Taking Mole Webquest Answers
  - Interactive Elements Mole Webguest Answers
- 8. Staying Engaged with Mole Webquest Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Mole Webquest Answers
- 9. Balancing eBooks and Physical Books Mole Webguest Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Mole Webquest Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Mole Webguest Answers
  - Setting Reading Goals Mole Webquest Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Mole Webguest Answers
  - Fact-Checking eBook Content of Mole Webquest Answers
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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