

# CHAPTER 10 The Mole

## Section 1 Measuring Matter

page 320–324

### Practice Problems

pages 323–324

1. Zinc (Zn) is used to form a corrosion-inhibiting surface on galvanized steel. Determine the number of Zn atoms in 2.50 mol of Zn.

$$2.50 \text{ mol Zn} \times \frac{6.02 \times 10^{23} \text{ atoms Zn}}{1 \text{ mol Zn}} \\ = 1.51 \times 10^{24} \text{ atoms of Zn}$$

2. Calculate the number of molecules in 11.5 mol of water (H<sub>2</sub>O).

$$11.5 \text{ mol H}_2\text{O} \times \frac{6.02 \times 10^{23} \text{ molecules H}_2\text{O}}{1 \text{ mol H}_2\text{O}} \\ = 6.92 \times 10^{24} \text{ molecules H}_2\text{O}$$

3. Silver nitrate (AgNO<sub>3</sub>) is used to make several different silver halides used in photographic films. How many formula units of AgNO<sub>3</sub> are there in 3.25 mol AgNO<sub>3</sub>?

$$3.25 \text{ mol AgNO}_3 \times \frac{6.02 \times 10^{23} \text{ formula units AgNO}_3}{1 \text{ mol AgNO}_3} \\ = 1.96 \times 10^{24} \text{ formula units AgNO}_3$$

4. **Challenge** Calculate the number of oxygen atoms in 5.00 mol of oxygen molecules. Oxygen is a diatomic molecule, O<sub>2</sub>.

$$5.00 \text{ mol O}_2 \times \frac{6.02 \times 10^{23} \text{ molecules O}_2}{1 \text{ mol O}_2}$$

$$\times \frac{2 \text{ O atoms}}{\text{molecule O}_2} = 6.02 \cdot 10^{24} \text{ atoms O}$$

5. How many moles contain each of the following?

- a.  $5.75 \times 10^{24}$  atoms Al

$$5.75 \times 10^{24} \text{ atoms Al} \times \frac{1 \text{ mol Al}}{6.02 \times 10^{23} \text{ atoms Al}} \\ = 9.55 \text{ mol Al}$$

- b.  $2.50 \times 10^{23}$  atoms Fe

$$2.50 \times 10^{23} \text{ atoms Fe} \times \frac{1 \text{ mol Fe}}{6.02 \times 10^{23} \text{ atoms Fe}} \\ = 4.15 \times 10^{-1} \text{ mol Fe}$$

6. **Challenge** Identify the representative particle for each formula and convert the given number of representative particles to moles.

- a.  $3.75 \times 10^{24}$  CO<sub>2</sub>

The representative particle is a molecule.

$$3.75 \times 10^{24} \text{ molecules CO}_2 \\ \times \frac{1 \text{ mol CO}_2}{6.02 \times 10^{23} \text{ molecules CO}_2} = 6.23 \text{ mol CO}_2$$

- b.  $3.58 \times 10^{23}$  ZnCl<sub>2</sub>

The representative particle is a formula unit.

$$3.58 \times 10^{23} \text{ formula units ZnCl}_2 \\ \times \frac{1 \text{ mol ZnCl}_2}{6.02 \times 10^{23} \text{ formula units ZnCl}_2} \\ = 0.595 \text{ mol ZnCl}_2$$

## Section 1 Review

page 324

7. **Explain** why chemists use the mole.

Chemists use the mole because it is a convenient way of knowing how many representative particles are in a sample.

8. **State** the mathematical relationship between Avogadro's number and 1 mol.

One mole contains Avogadro's number ( $6.02 \times 10^{23}$ ) of representative particles.

# Mole Solution Manual

**Richard Langley**



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### **Table of Contents Mole Solution Manual**

1. Understanding the eBook Mole Solution Manual
  - The Rise of Digital Reading Mole Solution Manual
  - Advantages of eBooks Over Traditional Books
2. Identifying Mole Solution Manual
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Mole Solution Manual
  - User-Friendly Interface
4. Exploring eBook Recommendations from Mole Solution Manual
  - Personalized Recommendations

- Mole Solution Manual User Reviews and Ratings
- Mole Solution Manual and Bestseller Lists
- 5. Accessing Mole Solution Manual Free and Paid eBooks
  - Mole Solution Manual Public Domain eBooks
  - Mole Solution Manual eBook Subscription Services
  - Mole Solution Manual Budget-Friendly Options
- 6. Navigating Mole Solution Manual eBook Formats
  - ePub, PDF, MOBI, and More
  - Mole Solution Manual Compatibility with Devices
  - Mole Solution Manual Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Mole Solution Manual
  - Highlighting and Note-Taking Mole Solution Manual
  - Interactive Elements Mole Solution Manual
- 8. Staying Engaged with Mole Solution Manual
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Mole Solution Manual
- 9. Balancing eBooks and Physical Books Mole Solution Manual
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Mole Solution Manual
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Mole Solution Manual
  - Setting Reading Goals Mole Solution Manual
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Mole Solution Manual
  - Fact-Checking eBook Content of Mole Solution Manual



- Distinguishing Credible Sources
13. Promoting Lifelong Learning
    - Utilizing eBooks for Skill Development
    - Exploring Educational eBooks
  14. Embracing eBook Trends
    - Integration of Multimedia Elements
    - Interactive and Gamified eBooks

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