

(c) Mass of 1 mole of anhydrous calcium chloride ( $\text{CaCl}_2$ ) =  $40 + 71 = 111 \text{ g}$

But 1 molecule of  $\text{CaCl}_2$  contains 2 chloride ions<sub>23</sub>  
 $\therefore 111 \text{ g}$  of anhydrous  $\text{CaCl}_2$  contains  $2 \times 6.023 \times 10^{23}$   
 $= 12 \times 10^{23}$  chloride ions

Q2) How many moles of  $\text{NaOH}$  are present in  $160 \text{ g}$  of it?

Soln: Mass of 1 mole of  $\text{NaOH} = 23 + 16 + 1 = 40 \text{ g}$

$40 \text{ g}$  of  $\text{NaOH}$  contains  $6 \times 10^{23}$  no. of molecules

Moles in  $160 \text{ g}$  of  $\text{NaOH} = \frac{160}{40} = \underline{4 \text{ moles}}$

Q3) What is the mass in grams of 2.5 moles of slaked lime  $\text{Ca(OH)}_2$ .

Soln: Mass of 1 mole of slaked lime  $\text{Ca(OH)}_2 = 40 + 32 + 2$   
 $= 74 \text{ g}$

So mass of 2.5 moles of  $\text{Ca(OH)}_2 = 2.5 \times 74 = \underline{185 \text{ g}}$

Means 2.5 moles of  $\text{Ca(OH)}_2$  weigh 185 g.

# Mole Concept Problems With Solutions

**Disha Experts**



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