

Molarity Problems

KEY

NOT IN ORDER

1. Sea water contains roughly 28.0 g of NaCl per liter. What is the molarity of sodium chloride in sea water?

$$\frac{28.0g \text{ NaCl}}{1} \times \frac{1 \text{ mol NaCl}}{58.44g} = \boxed{0.48 \text{ M}}$$

2. What is the molarity of 245.0 g of H_2SO_4 dissolved in 1.00 L of solution?

$$\frac{245.0g \text{ H}_2\text{SO}_4}{1} \times \frac{1 \text{ mol H}_2\text{SO}_4}{98.08g} = \boxed{2.5 \text{ M}}$$

3. What is the molarity of 5.30 g of Na_2CO_3 dissolved in 400.0 mL solution?

$$\frac{5.30g \text{ Na}_2\text{CO}_3}{1} \times \frac{1 \text{ mol Na}_2\text{CO}_3}{106g} = \frac{0.05 \text{ mol}}{.400 \text{ L}} = \boxed{0.125 \text{ M}}$$

4. What is the molarity of 5.00 g of NaOH in 750.0 mL of solution?

$$\frac{5.00g \text{ NaOH}}{1} \times \frac{1 \text{ mol NaOH}}{40g} = \frac{.125 \text{ mol}}{.750 \text{ L}} = \boxed{0.18 \text{ M}}$$

5. How many moles of Na_2CO_3 are there in 10.0 L of 2.0 M solution?

$$\frac{2.0 \text{ mol}}{1} \times \frac{10.0 \text{ L}}{1} = \boxed{20 \text{ mol Na}_2\text{CO}_3}$$

6. How many moles of Na_2CO_3 are in 10.0 mL of a 2.0 M solution?

$$\frac{2.0 \text{ mol}}{1} \times \frac{.01 \text{ L}}{1} = \boxed{0.02 \text{ mol Na}_2\text{CO}_3}$$

7. How many moles of NaCl are contained in 100.0 mL of a 0.20 M solution?

$$\frac{.20 \text{ mol}}{1} \times \frac{.100 \text{ L}}{1} = \boxed{0.02 \text{ mole NaCl}}$$

8. What weight (in grams) of NaCl would be contained in problem 7?

$$\frac{.02 \text{ mol NaCl}}{1} \times \frac{58.44 \text{ g NaCl}}{1 \text{ mol NaCl}} = \boxed{1.16 \text{ g NaCl}}$$

9. What weight (in grams) of H_2SO_4 would be needed to make 750.0 mL of 2.00 M solution?

$$\frac{2.00 \text{ mol}}{1} \times \frac{.750 \text{ L}}{1} \times \frac{98.08 \text{ g H}_2\text{SO}_4}{1 \text{ mol H}_2\text{SO}_4} = \boxed{147 \text{ g H}_2\text{SO}_4}$$

10. What volume (in mL) of 18.0 M H_2SO_4 is needed to contain 2.45 g H_2SO_4 ?

$$\frac{2.45g \text{ H}_2\text{SO}_4}{1} \times \frac{1 \text{ mol H}_2\text{SO}_4}{98.08g} \times \frac{1 \text{ L}}{18.0 \text{ mol H}_2\text{SO}_4} \times \frac{1000 \text{ mL}}{1 \text{ L}} = \boxed{1.39 \text{ mL}}$$

11. What volume (in mL) of 12.0 M HCl is needed to contain 3.00 moles of HCl?

$$\frac{3.00 \text{ mol HCl}}{1} \times \frac{1 \text{ L HCl}}{12.0 \text{ mol HCl}} \times \frac{1000 \text{ mL HCl}}{1 \text{ L HCl}} = \boxed{250 \text{ mL HCl}}$$

12. How many grams of $\text{Ca}(\text{OH})_2$ are needed to make 100.0 mL of 0.250 M solution?

$$\frac{.250 \text{ mol Ca}(\text{OH})_2}{1} \times \frac{.100 \text{ L}}{1} \times \frac{74 \text{ g Ca}(\text{OH})_2}{1 \text{ mol Ca}(\text{OH})_2} = \boxed{1.85 \text{ g Ca}(\text{OH})_2}$$

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Table of Contents Molarity Pogil Lemonade Answer Key

1. Understanding the eBook Molarity Pogil Lemonade Answer Key
 - The Rise of Digital Reading Molarity Pogil Lemonade Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Molarity Pogil Lemonade Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Molarity Pogil Lemonade Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Molarity Pogil Lemonade Answer Key
 - Personalized Recommendations
 - Molarity Pogil Lemonade Answer Key User Reviews and Ratings
 - Molarity Pogil Lemonade Answer Key and Bestseller Lists
5. Accessing Molarity Pogil Lemonade Answer Key Free and Paid eBooks
 - Molarity Pogil Lemonade Answer Key Public Domain eBooks
 - Molarity Pogil Lemonade Answer Key eBook Subscription Services
 - Molarity Pogil Lemonade Answer Key Budget-Friendly Options
6. Navigating Molarity Pogil Lemonade Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Molarity Pogil Lemonade Answer Key Compatibility with Devices
 - Molarity Pogil Lemonade Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Molarity Pogil Lemonade Answer Key
 - Highlighting and Note-Taking Molarity Pogil Lemonade Answer Key
 - Interactive Elements Molarity Pogil Lemonade Answer Key
8. Staying Engaged with Molarity Pogil Lemonade Answer Key

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Molarity Pogil Lemonade Answer Key
- 9. Balancing eBooks and Physical Books Molarity Pogil Lemonade Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Molarity Pogil Lemonade Answer Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Molarity Pogil Lemonade Answer Key
 - Setting Reading Goals Molarity Pogil Lemonade Answer Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Molarity Pogil Lemonade Answer Key
 - Fact-Checking eBook Content of Molarity Pogil Lemonade Answer Key
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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