

# MOLARITY BY DILUTION

79 + 52 = 131 Name \_\_\_\_\_

Acids are usually acquired from chemical supply houses in concentrated form. These acids are diluted to the desired concentration by adding water. Since moles of acid before dilution = moles of acid after dilution, and moles of acid =  $M \times V$  then,  
 $M_1 \times V_1 = M_2 \times V_2$  Solve the following problems.

1. How much concentrated 18 M sulfuric acid is needed to prepare 250 mL of a 6.0 M solution?

$$V_1 = \frac{M_2 V_2}{M_1} = \frac{(6.0 M)(250 \text{ mL})}{18 M}$$

83 mL

2. How much concentrated 12 M hydrochloric acid is needed to prepare 100 mL of a 2.0 M solution?

$$V_1 = \frac{M_2 V_2}{M_1} = \frac{(2.0 M)(100 \text{ mL})}{12 M}$$

17 mL

3. To what volume should 25 mL of 15 M nitric acid be diluted to prepare a 3.0 M solution?

$$V_2 = \frac{M_1 V_1}{M_2} = \frac{(15 M)(25 \text{ mL})}{3.0 M}$$

125 mL

4. To how much water should 50. mL of 12 M hydrochloric acid be added to produce a 4.0 M solution?

$$V_2 = \frac{M_1 V_1}{M_2} = \frac{(12 M)(50 \text{ mL})}{4.0 M} = 150 \text{ mL} - 50 \text{ mL}$$

100 mL H<sub>2</sub>O

5. To how much water should 100. mL of 18 M sulfuric acid be added to prepare a 1.5 M solution?

$$V_2 = \frac{M_1 V_1}{M_2} = \frac{(18 M)(100 \text{ mL})}{1.5 M} = 1200 \text{ mL} - 100 \text{ mL} = 1100 \text{ mL}$$

1.1 L

# Molarity Worksheet Answers Instructional Fair

**Instructional Fair (Firm)**



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## **Table of Contents Molarity Worksheet Answers Instructional Fair**

1. Understanding the eBook Molarity Worksheet Answers Instructional Fair
  - The Rise of Digital Reading Molarity Worksheet Answers Instructional Fair
  - Advantages of eBooks Over Traditional Books
2. Identifying Molarity Worksheet Answers Instructional Fair
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Molarity Worksheet Answers Instructional Fair
  - User-Friendly Interface
4. Exploring eBook Recommendations from Molarity Worksheet Answers Instructional Fair
  - Personalized Recommendations
  - Molarity Worksheet Answers Instructional Fair User Reviews and Ratings
  - Molarity Worksheet Answers Instructional Fair and Bestseller Lists

5. Accessing Molarity Worksheet Answers Instructional Fair Free and Paid eBooks
  - Molarity Worksheet Answers Instructional Fair Public Domain eBooks
  - Molarity Worksheet Answers Instructional Fair eBook Subscription Services
  - Molarity Worksheet Answers Instructional Fair Budget-Friendly Options
6. Navigating Molarity Worksheet Answers Instructional Fair eBook Formats
  - ePub, PDF, MOBI, and More
  - Molarity Worksheet Answers Instructional Fair Compatibility with Devices
  - Molarity Worksheet Answers Instructional Fair Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Molarity Worksheet Answers Instructional Fair
  - Highlighting and Note-Taking Molarity Worksheet Answers Instructional Fair
  - Interactive Elements Molarity Worksheet Answers Instructional Fair
8. Staying Engaged with Molarity Worksheet Answers Instructional Fair
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Molarity Worksheet Answers Instructional Fair
9. Balancing eBooks and Physical Books Molarity Worksheet Answers Instructional Fair
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Molarity Worksheet Answers Instructional Fair
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Molarity Worksheet Answers Instructional Fair
  - Setting Reading Goals Molarity Worksheet Answers Instructional Fair
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Molarity Worksheet Answers Instructional Fair
  - Fact-Checking eBook Content of Molarity Worksheet Answers Instructional Fair
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

#### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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