## The Molar Volume of a Gas

Computer 5

In this experiment, you will determine the molar volume of a gas by conducting a chemical reaction that produces a gas, as shown in the reaction equation below.

$$Mg(s) + 2HCl(aq) \rightarrow MgCl_2(aq) + H_2(g)$$

You will react a known mass of solid magnesium with an excess of hydrochloric acid, in a sealed vessel, and use the pressure change to calculate molar volume at STP.

#### OBJECTIVES

In this experiment, you will

- Measure the gas production of a chemical reaction by a pressure change.
- Determine the molar volume of the gas produced in the reaction.
- · Calculate the molar volume of a gas at STP.



Figure 1

#### MATERIALS

Vernier computer interface computer Vernier Gas Pressure Sensor Temperature Probe 600 mL or one liter beaker 10 mL graduated cylinder magnesium ribbon, pre-cut 3.0 M HCl solution small beaker for HCl solution ruler 125 mL Erlenmeyer flask syringe

# **Molar Volume Of Hydrogen Gas Lab Answers**

**L** Cohen

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U.S. Government Research & Development Reports ,1969-04 Exploring General Chemistry in the Laboratory Colleen F. Craig, Kim N. Gunnerson,2017-02-01 This laboratory manual is intended for a two semester general chemistry course The procedures are written with the goal of simplifying a complicated and often challenging subject for students by applying concepts to everyday life This lab manual covers topics such as composition of compounds reactivity stoichiometry limiting reactants gas laws calorimetry periodic trends molecular structure spectroscopy kinetics equilibria thermodynamics electrochemistry intermolecular forces solutions and coordination complexes By the end of this course you should have a solid understanding of the basic concepts of chemistry which will give you confidence as you embark on your career in science

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