Colligative properties are properties that depend only upon the number of solute atoms, ions, or molecules in a solution and not on the nature of those atoms, ions or molecules. Freezing point depression and boiling point elevation are examples of colligative properties. Raoult discovered that the addition of solute particles causes the boiling point of a solution to be elevated and the freezing point to be depressed. It has been found that I mole of nonvolatile, nonionizing solute particles will raise the boiling point of 1.0 kg of water by 0.52 °C, to 100.52 °C at 1 atm of pressure. The same concentration will lower the freezing point of 1.0 kg of water by 1.86 °C, to -1.86 °C. These two numbers apply only to water and are called the molal boiling point and freezing point constants for water. Other solvents have different constants.

PROPERTY	BOILING POINT ELEVATION	FREEZING POINT DEPRESSION
EQUATION	Boiling Point Elevation: A solution will boil at a higher temperature than the pur solvent.	
	$\Delta T_b = i K_b m$	$\Delta T_f = i K_f m$
	$\Delta T_b$ = change in boiling temp. in (°C)	$\Delta T_f$ = change in freezing temp. in (°C)
	i = a unitless constant (called the van't Hoff factor) associated with the degree of dissociation of the solute in a solvent.	i = a unitless constant (called the van't Hoff factor)associated with the degree of dissociation of the solute in a solvent.
	K <sub>b</sub> = molal boiling point elevation constant (°C/m)	K <sub>f</sub> = molal freezing point lowering constant (°C/m)
	For H <sub>2</sub> O: $K_b = .52^{\circ}C/m$	For H <sub>2</sub> O: K <sub>f</sub> = 1.86°C/m
	m = molality of the solution	m = molality of the solution
GUIDLINES	i =  1 for substances that do not ionize 2 for substances which ionize into two ions like LiCl 3 for substances which ionize into three ions like CaCl <sub>2</sub> etc.	i =  1 for substances that do not ionize 2 for substances which ionize into two ions like LiCl 3 for substances which ionize into three ions like CaCl <sub>2</sub> etc.
	i may have to be calculated if a %-of dissociation is given for a solute.	i may have to be calculated if a % of dissociation is given for a solute.
	If the solvent is not water you must find its K <sub>b</sub> .	If the solvent is not water you must finits K <sub>f</sub> .

#### EXAMPLE #1

If 90.0 g of nonionizing C<sub>6</sub>H<sub>1</sub>2O<sub>6</sub>, are dissolved in 255 g of H<sub>2</sub>O, what will be the boiling point of the resulting solution?

First calculate the molality (m) of the solution:

$$m = \frac{92g}{(180 \frac{g}{\text{mole}})(0.255 \text{ kg of solvent})} = 2.0 \text{ m}$$

Next apply the boiling point elevation equation: [i = 1 for a nonionizing solute.]  $\Delta T_b = i K_b m$   $\Delta T_b = (1)(0.52^{\circ}C/m)(2.0 m)$   $\Delta T_b = 1.04 ^{\circ}C$ BP =  $(\Delta T_b + 100.00 ^{\circ}C)$ BP =  $(1.04^{\circ}C + 100.00 ^{\circ}C) = 101.04^{\circ}C$ 

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- 1. Understanding the eBook Modern Chemistry Study Guide Colligative Properties Answers
  - The Rise of Digital Reading Modern Chemistry Study Guide Colligative Properties Answers
  - Advantages of eBooks Over Traditional Books
- 2. Identifying Modern Chemistry Study Guide Colligative Properties Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Modern Chemistry Study Guide Colligative Properties Answers
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from Modern Chemistry Study Guide Colligative Properties Answers
  - Personalized Recommendations
  - Modern Chemistry Study Guide Colligative Properties Answers User Reviews and Ratings
  - Modern Chemistry Study Guide Colligative Properties Answers and Bestseller Lists
- 5. Accessing Modern Chemistry Study Guide Colligative Properties Answers Free and Paid eBooks
  - Modern Chemistry Study Guide Colligative Properties Answers Public Domain eBooks
  - Modern Chemistry Study Guide Colligative Properties Answers eBook Subscription Services
  - Modern Chemistry Study Guide Colligative Properties Answers Budget-Friendly Options
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- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Modern Chemistry Study Guide Colligative Properties Answers
  - Highlighting and Note-Taking Modern Chemistry Study Guide Colligative Properties Answers
  - Interactive Elements Modern Chemistry Study Guide Colligative Properties Answers
- 8. Staying Engaged with Modern Chemistry Study Guide Colligative Properties Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Modern Chemistry Study Guide Colligative Properties Answers
- 9. Balancing eBooks and Physical Books Modern Chemistry Study Guide Colligative Properties Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Modern Chemistry Study Guide Colligative Properties Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Modern Chemistry Study Guide Colligative Properties Answers
  - Setting Reading Goals Modern Chemistry Study Guide Colligative Properties Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Modern Chemistry Study Guide Colligative Properties Answers
  - $\circ \ \ \text{Fact-Checking eBook Content of Modern Chemistry Study Guide Colligative Properties Answers} \\$
  - Distinguishing Credible Sources
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