

## Stoichiometry Reflect #2

Learning Target... I can use stoichiometry to calculate amounts

1a. Balance this equation



1b. Determine how many moles of ammonium oxide was used if the student weighed out 55.2 g of it.

$$55.2 \text{ g (NH}_4\text{)}_2\text{O} \left( \frac{1 \text{ mol A}}{52.10 \text{ g}} \right) = 1.06 \text{ moles (NH}_4\text{)}_2\text{O}$$

1c. Determine the amount of moles of ammonium nitride produced in this reaction.

$$1.06 \text{ moles A} \left( \frac{2 \text{ mol B}}{3 \text{ mol A}} \right) = 0.706 \text{ moles (NH}_4\text{)}_3\text{N}$$

1d. Determine how many grams of ammonium nitride produced in this reaction.

$$0.706 \text{ moles} \left( \frac{108.11 \text{ g}}{1 \text{ mol B}} \right) = 48.1 \text{ g (NH}_4\text{)}_3\text{N}$$

2. Using the following chemical reaction:

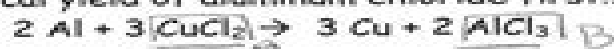


and, knowing that 61.6 g of copper was produced in this experiment, how many molecules of mercury nitrate was also produced?

$$61.6 \text{ g Cu} \left( \frac{1 \text{ mol A}}{63.55 \text{ g}} \right) \left( \frac{2 \text{ mol B}}{1 \text{ mol A}} \right) \left( \frac{6.02 \times 10^{23} \text{ mc}}{1 \text{ mol B}} \right) = 1.17 \times 10^{24} \text{ mc HgNO}_3$$

3. Challenge...

What is the percent yield of the following reaction if 12.8 g of copper (II) chloride was used in the lab, and 7.75 g of aluminum chloride was measured at the end of the experiment? (hint: you must calculate the theoretical yield of aluminum chloride first!!)



$$12.8 \text{ g A} \left( \frac{1 \text{ mol A}}{134.45 \text{ g}} \right) \left( \frac{2 \text{ mol B}}{3 \text{ mol A}} \right) \left( \frac{133.33 \text{ g}}{1 \text{ mol B}} \right) = 8.46 \text{ g AlCl}_3$$

theoretical

$$\% \text{ yield} = \frac{7.75 \text{ g}}{8.46 \text{ g}} = 91.4\%$$

Show work on back

# Modern Chemistry Stoichiometry Test Answers

**Roman Wölfel**



## **Modern Chemistry Stoichiometry Test Answers:**

## Modern Chemistry Stoichiometry Test Answers Book Review: Unveiling the Power of Words

In a world driven by information and connectivity, the power of words has become more evident than ever. They have the capability to inspire, provoke, and ignite change. Such could be the essence of the book **Modern Chemistry Stoichiometry Test Answers**, a literary masterpiece that delves deep into the significance of words and their affect our lives. Compiled by a renowned author, this captivating work takes readers on a transformative journey, unraveling the secrets and potential behind every word. In this review, we shall explore the book's key themes, examine its writing style, and analyze its overall impact on readers.

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