

Calculate the partial pressure of nitrogen gas, the moles of dry gas present and the total pressure in the system. Show your work.

$$\begin{aligned}
 PV &= nRT \rightarrow P_{\text{H}_2} = nRT/V \rightarrow (0.0540 \text{ mol} \times 0.08206 \text{ L atm/K mol} \times 300 \text{ K})/2.0 \text{ L} = \\
 P_{\text{H}_2} &= 0.632 \text{ atm} \\
 &\rightarrow n_{\text{H}_2} = PV/RT \rightarrow (0.91 \text{ atm} \times 2.0 \text{ L})/(0.08206 \text{ L atm/K mol} \times 300 \text{ K}) = \\
 n_{\text{H}_2} &= 0.0720 \text{ mol} \\
 P_{\text{tot}} &= P_{\text{H}_2} + P_{\text{N}_2} \\
 &= 0.91 \text{ atm} + 0.632 \text{ atm} = P_{\text{tot}} = 1.54 \text{ atm}
 \end{aligned}$$

7. Calculate the volume of hydrogen gas produced at 4.50 atm and -30°C by the reaction of 26.5g of calcium with excess hydrobromic acid according to the unbalanced equation. Show your work.



& 244K

$$PV = nRT \rightarrow 26.5 \text{ g Ca} \times (1 \text{ mol Ca}/40.08 \text{ g Ca}) \times (1 \text{ mol H}_2/1 \text{ mol Ca}) = 0.661 \text{ mol H}_2$$

$$V = nRT/P = (0.661 \text{ mol} \times 0.08206 \text{ L atm/K mol} \times 244 \text{ K})/4.50 \text{ atm} = V = 2.94 \text{ L}$$

8. Hydrogen Sulfide reacts with sulfur dioxide to produce water and sulfur. If 3.2g of sulfur was produced under the following conditions: 6.0 L @ 750 torr, what is the temperature of the hydrogen sulfide gas, in °C?



$$3.2 \text{ g S} \times (1 \text{ mol S}/32.07 \text{ g S}) \times (2 \text{ mol H}_2\text{S}/3 \text{ mol S}) = 0.067 \text{ mol H}_2\text{S}$$

$$PV = nRT \rightarrow T = PV/nR \rightarrow (0.987 \text{ atm} \times 6.0 \text{ L})/(0.067 \text{ mol} \times 0.08206 \text{ L atm/K mol}) = 1077.12 \text{ K}$$

$$= 804^\circ\text{C}$$

SLT - Student will apply Kinetic Molecular Theory when describing a novel, real-world situation in terms of collisions and motions of particles.

9. In the summer, the tire pressure sensor in your car alerts you that your tires need to be checked. At the gas station, you find that the tire pressure is too high inside the tire. However, you have not added additional gas. Using the Kinetic Molecular Theory, explain why this occurs, making sure to refer to the motion of the gas molecules.

Since the summer is known for high temps, a high T would cause for the motion of gas molecules inside the tire to increase. With the increase in motion comes an increase in the number of collisions. With more collisions occurring, this results in an increase in pressure inside the tire.

Modern Chemistry Gases Review Answer Key

M Carnoy



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