

Chemistry 11 Stoichiometry Review

Chapter 2 and 4 Questions

1. Iron metal is produced from the mineral hematite (Fe_2O_3) and pure carbon, as given by the equation:



Balance this equation. Then, use the mole ratios therein to perform the calculations below.

- How many moles of C are required to completely react with 14.251 moles of Fe_2O_3 (hematite)?
- How many moles of Fe_2O_3 must react to generate 0.986 moles of pure iron?
- How many moles of CO_2 will be generated if 2.774 moles of carbon completely reacts?

2. When nitroglycerin explodes, the following reaction occurs:



If 0.398 moles of nitroglycerine explodes, how many moles of each product would be formed?

3. The element phosphorus naturally occurs as $\text{P}_4 (s)$. It reacts readily with magnesium according to the following equation:



What mass (in g) of P_4 is required to completely react with 73.4 grams of Mg?

4. Ammonia reacts with oxygen according to the following equation:



How many *molecules* of water will be generated by the complete reaction of 191.4 grams of ammonia?

5. Carbon tetrachloride may be produced via the reaction between methane and chlorine gases:



If this reaction generates 2.50L of CCl_4 , how many grams of HCl were also generated?

6. Aluminum reacts with bromine to produce aluminum bromide according to the following equation:



In the lab you perform this reaction using an excess of aluminum and 55.7 grams of bromine. You end up collecting 60.0 grams of product aluminum bromide. What is your percent yield?

7. Iron reacts with oxygen to give iron(III) oxide according to the following balanced equation:



Identify the limiting reactant in each of the following reactant mixtures:

- 0.288 moles Fe and 0.240 moles O_2
- 10.0 grams Fe and 10.0 grams O_2
- 1.45×10^{24} atoms of Fe and 7.21×10^{23} molecules of O_2

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