

Section 1 Describing Chemical Reactions Pages 261-275

1. Give four indications of a chemical reaction.
2. How do you prove a reaction (chemical change) has taken place?
3. Give three characteristics of a chemical equation.
4. List the diatomic elements.
5. List the molecular elements.
6. Define coefficients.
7. Define word equation and give an example.
8. Define a formula equation and give an example.
9. List and explain the state of matter symbols. Where are they found in the chemical equation?
10. Find the symbols for the following: yields, reversible reaction, precipitate, gas, heating reactants, specific pressure, specific temperature, catalyst, electrolysis.
11. If balancing an equation was a game or puzzle, what would the goal of the game be?
12. If balancing an equation was a game or puzzle, what would the rules of the game be?
13. On page 271 there are some helpful tips for balancing equations. List as many as you can find.
14. What are the possible units for coefficients?
15. Is the ratio of the coefficients of the reactants the same as the ratio of the grams of the reactants?

Section 2 Types of Chemical Reactions Pages 276-284

16. For each type of reactions list the following.
 - a. A definition
 - b. A generic equation using the letters A, X, B or Y.
 - c. The kinds of reactions belonging to that type.
 - d. An example equation for each kind of reaction.

Section 3 Activity Series of the Elements Pages 285-288

17. Define the activity series of the elements.
18. Give a few of the elements in the activity series that are the most reactive.
19. Give a few of the elements in the activity series that are the least reactive.
20. How is the activity series used to determine if a single displacement reaction will occur?

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