

Acids and bases review sheet Answer key

1. List the differences between an acid and a base. Your description should include such things as whether both are electrolytes, which contains a hydrogen ion and which includes a hydroxide ion, and you should also include two examples of an acid and two examples of a base that are present in your household.

Acid	base
Conducts electricity	conducts electricity
Contains hydrogen ion	contains hydroxide ion
Proton donor	proton acceptor
Turns litmus paper red	turns litmus paper blue
React with most metals	releases hydroxide ion
Releases hydronium	Ex: Windex,
Ex: lemon	bleach
Vinegar	soap
Battery acid (sulfuric acid)	

2. Listed below are some of the properties of acids and bases. Fill in the blanks with the appropriate word, **acids** or **bases**:

- A** produce hydrogen ions (H^+) in solution
- B** have a bitter taste
- B** produce hydroxide ions (OH^-) ions
- A** react with bases to form salts
- A** have a sour taste
- B** have a slippery, soapy feel
- B** react with acids to form salts
- A** react with many metals

3. Write an equation for the dissociation of the following **acids** in water; follow example

- $HClO_4 \rightarrow H^+ + ClO_4^-$
- $H_2SO_4 \rightarrow 2H^+ + SO_4^{2-}$
- $HC_2H_3O_2 \rightarrow H^+ + C_2H_3O_2^-$
- $H_2S \rightarrow 2H^+ + S^{2-}$
- $HCl \rightarrow H^+ + Cl^-$
- $HNO_3 \rightarrow H^+ + NO_3^-$

4. Write an equation for the dissociation (ionization) of the following **bases** in water: (follow example)

- $Mg(OH)_2 \rightarrow Mg^{2+} + 2OH^-$
- $NaOH \rightarrow Na^+ + OH^-$
- $Ba(OH)_2 \rightarrow Ba^{2+} + 2(OH^-)$
- $KOH \rightarrow K^+ + OH^-$
- $Mg(OH)_2 \rightarrow Mg^{2+} + 2(OH^-)$
- $LiOH \rightarrow Li^+ + OH^-$

5) A. What type of reaction is a **neutralization** reaction? **Double replacement**

B. Write the net equation for neutralization:



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G Thomas



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Table of Contents Modern Chemistry Acids And Bases Review Answers

1. Understanding the eBook Modern Chemistry Acids And Bases Review Answers
 - The Rise of Digital Reading Modern Chemistry Acids And Bases Review Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Modern Chemistry Acids And Bases Review Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Modern Chemistry Acids And Bases Review Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Modern Chemistry Acids And Bases Review Answers
 - Personalized Recommendations
 - Modern Chemistry Acids And Bases Review Answers User Reviews and Ratings
 - Modern Chemistry Acids And Bases Review Answers and Bestseller Lists

5. Accessing Modern Chemistry Acids And Bases Review Answers Free and Paid eBooks
 - Modern Chemistry Acids And Bases Review Answers Public Domain eBooks
 - Modern Chemistry Acids And Bases Review Answers eBook Subscription Services
 - Modern Chemistry Acids And Bases Review Answers Budget-Friendly Options
6. Navigating Modern Chemistry Acids And Bases Review Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Modern Chemistry Acids And Bases Review Answers Compatibility with Devices
 - Modern Chemistry Acids And Bases Review Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Modern Chemistry Acids And Bases Review Answers
 - Highlighting and Note-Taking Modern Chemistry Acids And Bases Review Answers
 - Interactive Elements Modern Chemistry Acids And Bases Review Answers
8. Staying Engaged with Modern Chemistry Acids And Bases Review Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Modern Chemistry Acids And Bases Review Answers
9. Balancing eBooks and Physical Books Modern Chemistry Acids And Bases Review Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Modern Chemistry Acids And Bases Review Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Modern Chemistry Acids And Bases Review Answers
 - Setting Reading Goals Modern Chemistry Acids And Bases Review Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Modern Chemistry Acids And Bases Review Answers
 - Fact-Checking eBook Content of Modern Chemistry Acids And Bases Review Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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