

# Chemistry Review Solutions

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No work shown = no credit. Please write legibly. Numerical answers must be rounded to the appropriate number of significant figures or decimal points and must contain the appropriate unit. Numerical answers that have the wrong number of significant figures, the wrong unit, or no unit will lose 20% credit.

- (9 points) Tell whether each stage of the solution process **absorbs** or **releases** energy. Each item is worth 3 points.  
(A) Solute particles separate.  
(B) Solvent particles separate.  
(C) Solvent particles surround solute particles (solvation).
- (6 points) The heats of solution of two solutes are given below. If you dissolve each substance in water, tell whether the surroundings should get warmer or cooler as the solute dissolves. Each item is worth 3 points.  
(A)  $\text{HBr}$ ,  $\Delta H_{\text{soln}} = -85.1 \text{ kJ/mol}$  (B)  $\text{Na}_2\text{C}_2\text{O}_4 \cdot 3\text{H}_2\text{O}$ ,  $\Delta H_{\text{soln}} = +19.7 \text{ kJ/mol}$
- (10 points) Use your solubility rules to determine whether each of the following compounds is soluble or insoluble in water. Write S for soluble and I for insoluble. Each item is worth 2 points.  
(A)  $\text{AgCl}$  (C)  $\text{KC}_2\text{H}_3\text{O}_2$  (E)  $\text{Zn}_3(\text{PO}_4)_2$   
(B)  $\text{CuBr}_2$  (D)  $\text{Ba}(\text{OH})_2$
- (25 points) Use your solubility curves graph to answer the following questions. Each item is worth 5 points.  
(A) How many grams of  $\text{KNO}_3$  can dissolve in 100 grams of water at  $25^\circ\text{C}$ ?  
(B) How many grams of  $\text{KI}$  can dissolve in 50 grams of water at  $10^\circ\text{C}$ ?  
(C) If you wish to dissolve 140 grams of  $\text{NaNO}_3$  in 100 grams of water, to what temperature should the water be heated?  
(D) A saturated solution of  $\text{KCl}$  is prepared in 100 grams of water at  $30^\circ\text{C}$ . If the solution is heated to  $45^\circ\text{C}$ , how many additional grams of  $\text{KCl}$  can be dissolved?  
(E) A saturated solution of  $\text{NH}_4\text{Cl}$  is prepared at  $70^\circ\text{C}$ . To what temperature should the solution be cooled to force 10 grams of  $\text{NH}_4\text{Cl}$  to precipitate?
- (5 points) What is the %w/w of a phenolphthalein solution if 0.500 grams of phenolphthalein are dissolved in 15.0 grams of ethanol?
- (5 points) What is the %v/v of an isopropanol solution if 90.0 mL of solution contains 45.0 mL of isopropanol?
- (10 points) What is the Molar concentration of a solution if 8.00 grams of  $\text{NaOH}$  (molar mass = 40.00 g/mol) are dissolved in enough water to make 100. mL of solution?

# Modern Chemistry Review Solutions Answers

**Wade A. Freeman**



## Modern Chemistry Review Solutions Answers:

Modern Chemistry Holt Rinehart & Winston, Holt, Rinehart and Winston Staff, 2001      *Modern Chemistry* Nicholas D. Tzimopoulos, 1993      Mining and Chemical Engineering Review, 1920      Catalog of Copyright Entries. Third Series Library of Congress. Copyright Office, 1976      **Principles of Modern Chemistry** David W. Oxtoby, H. Pat Gillis, Laurie J. Butler, 2016-01-01 Long considered the standard for honors and high level mainstream general chemistry courses PRINCIPLES OF MODERN CHEMISTRY continues to set the standard as the most modern rigorous and chemically and mathematically accurate text on the market This authoritative text features an atoms first approach and thoroughly revised chapters on Quantum Mechanics and Molecular Structure Chapter 6 Electrochemistry Chapter 17 and Molecular Spectroscopy and Photochemistry Chapter 20 In addition the text utilizes mathematically accurate and artistic atomic and molecular orbital art and is student friendly without compromising its rigor End of chapter study aids focus on only the most important key objectives equations and concepts making it easier for students to locate chapter content while applications to a wide range of disciplines such as biology chemical engineering biochemistry and medicine deepen students understanding of the relevance of chemistry beyond the classroom      *Basic Concepts of Chemistry, Study Guide* Leo J. Malone, 2000 The text s three main goals are to introduce chemistry as a living relevant science to encourage learning and critical thinking and to help readers overcome the math difficulties that impede their progress in chemistry Designed to help readers master the principles of general chemistry As a prep book it promotes active involvement with the material There are special features throughout that reinforce concepts and help to develop strong problem solving and study skills Updated to Include an Interactive Learning Ware problems CD containing several of the chapter ending problems from the book in an interactive tutorial with feedback to help readers set up and solve problems      **The World Review**, 1927      Modern Chemistry H. Clark Metcalfe, Joseph F. Castka, 1966      Volume 1: Modern Electrochemistry John O'M. Bockris, Amulya K.N. Reddy, 1998-06-30 This book had its nucleus in some lectures given by one of us J O M B in a course on electrochemistry to students of energy conversion at the University of Pennsylvania It was there that he met a number of people trained in chemistry physics biology metallurgy and materials science all of whom wanted to know something about electrochemistry The concept of writing a book about electrochemistry which could be understood by people with very varied backgrounds was thereby engendered The lectures were recorded and written up by Dr Klaus Muller as a 293 page manuscript At a later stage A K N R joined the effort it was decided to make a fresh start and to write a much more comprehensive text Of methods for direct energy conversion the electrochemical one is the most advanced and seems the most likely to become of considerable practical importance Thus conversion to electrochemically powered transportation systems appears to be an important step by means of which the difficulties of air pollution and the effects of an increasing concentration in the atmosphere of carbon dioxide may be met Cor sion is recognized as having an electrochemical basis The synthesis of nylon

now contains an important electrochemical stage Some central biological mechanisms have been shown to take place by means of electrochemical reactions A number of American organizations have recently recommended greatly increased activity in training and research in electrochemistry at universities in the United States The Utah Educational Review, 1958 *Study and Problem Solving Guide to Accompany Principles of Modern Chemistry, Oxtoby/Nachtrieb* Wade A.

Freeman, 1986 **A Guide to Problems in Modern Electrochemistry 1** Maria E. Gamboa-Aldeco, Robert J.

Gale, 2011-06-27 It has been always an incentive for students to find whether his her efforts to solve exercises give correct results or to find tips for problems that he she finds more difficult These are the main reasons for the appearance of the present book As part of the textbook Modern Electrochemistry 1 Ionics A Guide to Problems in Modern Electrochemistry Part 1 Ionics compiles many of the solutions to the exercises and problems presented in the text as well as many new problems

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- Chapter 4: Modern Chemistry Review Solutions Answers in Specific Contexts
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