

Name _____ Date _____ Class _____

Modern Chemistry • CHAPTER 6**HOMEWORK 6-5***(pp. 172–173)***VOCABULARY**

Define.

1. double bond _____

2. a. triple bond _____
b. How is a triple bond shown? _____
3. resonance _____

4. resonance structures _____

GRAPHIC ORGANIZER

On a separate sheet of paper, use Table 6-2 on page 173 to create a graph comparing bond length to bond energy. Plot bond length on the x -axis (in pm) and bond energy on the y -axis (in kJ/mol). Depict all the bonds shown in the table on your graph.



What does your graph show about the relationship between bond length and bond energy? Which bonds are stronger—single, double, or triple?

STANDARDIZED TEST PREP

Circle the letter of the best answer.



1. The Lewis structure for C_2H_4 is $\text{H} \text{---} \text{C} = \text{C} \text{---} \text{H}$. This Lewis structure shows that
 - a. the bonds between carbon and hydrogen are double bonds.
 - b. each hydrogen atom has two valence electrons.
 - c. the bond between the carbon atoms is a double bond.
 - d. the bond energies of the carbon-hydrogen bonds are higher than the bond energy of the carbon-carbon bond.
2. Which is the correct Lewis structure for HCl?
 - a. $\text{H} \text{---} \ddot{\text{Cl}} \text{---}$
 - b. $\text{H} \text{---} \text{Cl}$
 - c. $\text{H} \text{---} \ddot{\text{Cl}} \text{---}$
 - d. $:\text{H}::\text{Cl}:$

Modern Chemistry

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