



- 2 D, E, F, and G are four consecutive elements in the fourth period of the Periodic Table. (The letters are not the actual symbols of the elements.)

D is a soft, silvery metal with a melting point just above room temperature. Its amphoteric oxide, D_2O_3 , has a melting point of 1900°C and can be formed by heating D in oxygen.

G is a solid that can exist as several different allotropes, most of which contain G_8 molecules. G burns in air to form GO_2 which dissolves in water to form an acidic solution. This solution reacts with sodium hydroxide to form the salt Na_2GO_3 .

- (a) Suggest the identities of D and G.

D G [1]

- (b) Write equations for the reactions of D_2O_3 with

(i) hydrochloric acid,

..... [2]

(ii) sodium hydroxide.

..... [2]

- (c) Suggest the type of bonding and structure in D_2O_3 .

..... [1]

- (d) Write an equation for the formation of an acidic solution when GO_2 dissolves in water.

..... [1]

[Total: 7]

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