

Section 1

Topic: Moles

Section 1: Introduction to Moles and Molar Mass

Section 1: Introduction to Moles

1. The mole is the SI unit for the amount of substance. It is defined as the amount of substance that contains exactly 6.022×10^{23} elementary particles (atoms, molecules, ions, etc.).
2. The mole is a unit of measurement for the amount of substance. It is defined as the amount of substance that contains exactly 6.022×10^{23} elementary particles (atoms, molecules, ions, etc.).
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Section 2: Molar Mass

The molar mass of a substance is the mass of one mole of that substance. It is expressed in g mol^{-1} . For example, the molar mass of water (H_2O) is 18 g mol^{-1} .

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Section 3: Molar Mass and the Molar Mass of a Compound

The molar mass of a compound is the sum of the molar masses of the elements that make up the compound. For example, the molar mass of water (H_2O) is 18 g mol^{-1} .

1. The molar mass of a compound is the sum of the molar masses of the elements that make up the compound. For example, the molar mass of water (H_2O) is 18 g mol^{-1} .
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Section 4: Molar Mass

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