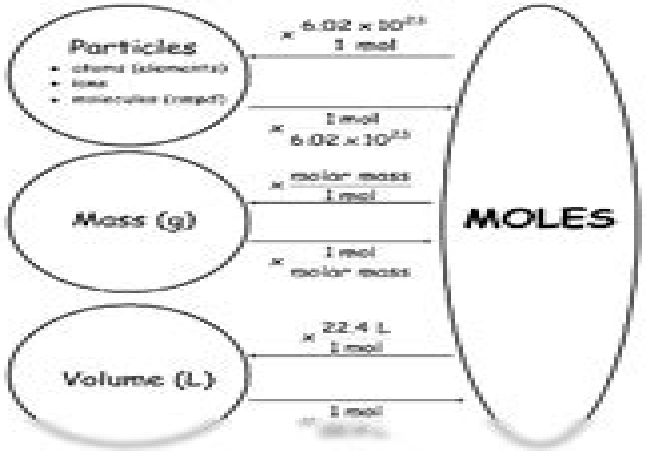
# Honors Chemistry B- Final Exam Formula Sheet

Note: You will not use **all** of these formulas for this test

**Temperature Formulas:**

$°C = (°F - 32) \cdot \frac{5}{9}$ $\quad °F = \frac{9}{5}°C + 32$ $\quad K = °C + 273$

## Mole Conversion Diagram




$$M_1 \times V_1 = M_2 \times V_2$$

$$\text{Percent by Mass} = \frac{\text{Mass of Solute}}{\text{Mass of Solution}} \times 100$$

$$\text{Percent by Volume} = \frac{\text{Volume of Solute}}{\text{Volume of Solution}} \times 100$$

Mole fractions: $\chi_A = \frac{n_A}{n_A + n_B}$ $\quad \chi_B = \frac{n_B}{n_A + n_B}$

1 Calorie = 1 kilocalorie = 1000 calories

$$Q = mc\Delta T$$

$$q_{sys} = \Delta H = -q_{surr} = -m \times C \times \Delta T$$

$$\Delta H^0 = \Delta H_f^0(\text{products}) - \Delta H_f^0(\text{reactants})$$

[illegible]

$$\Delta G = \Delta H - T\Delta S$$

[illegible]

$$K_w = [H^+][OH^-] = 1.0 \times 10^{-14}$$

$$pH = -\log[H^+]$$

$$pOH = -\log[OH^-]$$

$$K_a = \frac{[H^+][A^-]}{[HA]}$$

[illegible]

# High School Chemistry Honors Study Guide Final

Daniel F McAuley

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