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Chemistry
 Chemistry
 Chemistry

02 E I

Two hours
 Two hours
 Two hours

Instructions:

- Periodic Table is provided.
- This paper consists of 09 pages.
- Answer all the questions.
- Use of calculators is not allowed.
- Write your Index Number in the space provided in the answer sheet.
- Follow the instructions given on the back of the answer sheet carefully.
- In each of the questions 1 to 50, pick one of the alternatives from (1), (2), (3), (4), (5) which is correct or most appropriate and mark your response on the answer sheet with a cross (x) in accordance with the instructions given on the back of the answer sheet.

Universal gas constant $R = 8.314 \text{ J K}^{-1} \text{ mol}^{-1}$

Planck's constant $h = 6.626 \times 10^{-34} \text{ J s}$

Avogadro constant $N_A = 6.022 \times 10^{23} \text{ mol}^{-1}$

Velocity of light $c = 3 \times 10^8 \text{ m s}^{-1}$

- Which of the following electronic transitions belongs to the visible region of the line spectrum of atomic hydrogen? (n = Principal quantum number)
 - $n = 5 \rightarrow n = 3$
 - $n = 4 \rightarrow n = 2$
 - $n = 1 \rightarrow n = 2$
 - $n = 3 \rightarrow n = 1$
 - $n = 2 \rightarrow n = 1$
- Select the incorrect statement.
 - Pauli exclusion principle excludes the possibility of the presence of a third electron in an orbital.
 - In a potassium atom, the number of electrons having quantum numbers n (principal quantum number) = 3 and m_l (magnetic quantum number) = 0 is 4.
 - The effective nuclear charge felt by a valence electron in nitrogen (N), is greater than the effective nuclear charge felt by a valence electron in carbon (C).
 - Among the ions Na^+ , Mg^{2+} , K^+ and Ca^{2+} , the two ions closest to each other in size are K^+ and Mg^{2+} .
 - Electron gain energy of carbon is negative.
- The increasing order of the second ionization energy ($\text{X}^+(\text{g}) \rightarrow \text{X}^{2+}(\text{g}) + e$) of Be, B and O is,
 - $\text{Be} < \text{B} < \text{O}$
 - $\text{Be} < \text{O} < \text{B}$
 - $\text{B} < \text{O} < \text{Be}$
 - $\text{B} < \text{Be} < \text{O}$
 - $\text{O} < \text{Be} < \text{B}$
- The shapes of F_3ClO , FCIO_2 and FCIO_3 are respectively,
 - tetrahedral, trigonal planar and see-saw.
 - square planar, trigonal planar and tetrahedral.
 - see-saw, trigonal pyramidal and square planar.
 - tetrahedral, trigonal pyramidal and see-saw.
 - see-saw, trigonal pyramidal and tetrahedral.

[See page two]

Gde Chemistry 20memo

IM Harris



Gde Chemistry 20memo:

This book delves into Gde Chemistry 20memo. Gde Chemistry 20memo is an essential topic that needs to be grasped by everyone, from students and scholars to the general public. This book will furnish comprehensive and in-depth insights into Gde Chemistry 20memo, encompassing both the fundamentals and more intricate discussions.

1. The book is structured into several chapters, namely:

- Chapter 1: Introduction to Gde Chemistry 20memo
- Chapter 2: Essential Elements of Gde Chemistry 20memo
- Chapter 3: Gde Chemistry 20memo in Everyday Life
- Chapter 4: Gde Chemistry 20memo in Specific Contexts
- Chapter 5: Conclusion

2. In chapter 1, this book will provide an overview of Gde Chemistry 20memo. This chapter will explore what Gde Chemistry 20memo is, why Gde Chemistry 20memo is vital, and how to effectively learn about Gde Chemistry 20memo.

3. In chapter 2, the author will delve into the foundational concepts of Gde Chemistry 20memo. This chapter will elucidate the essential principles that need to be understood to grasp Gde Chemistry 20memo in its entirety.

4. In chapter 3, this book will examine the practical applications of Gde Chemistry 20memo in daily life. This chapter will showcase real-world examples of how Gde Chemistry 20memo can be effectively utilized in everyday scenarios.

5. In chapter 4, this book will scrutinize the relevance of Gde Chemistry 20memo in specific contexts. This chapter will explore how Gde Chemistry 20memo is applied in specialized fields, such as education, business, and technology.

6. In chapter 5, this book will draw a conclusion about Gde Chemistry 20memo. The final chapter will summarize the key points that have been discussed throughout the book.

The book is crafted in an easy-to-understand language and is complemented by engaging illustrations. It is highly recommended for anyone seeking to gain a comprehensive understanding of Gde Chemistry 20memo.

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