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පැප දෙකයි මුදුක්තු හත්බුණුගොඩ Two hours

Imstructionse

- * Periodic Table is provided.
- This paper consists of 09 pages.
- · Answer all the questions.
- Use of calculators is not allowed.
- Write your Index Number in the space provided in the answer sheet.
- Follow the instructions given on the back of the answer sheet carefully.
- * In each of the questions 1 to 50, pick one of the alternatives from (1), (2), (3), (4), (5) which is correct or most appropriate and mark your response on the answer sheet with a cross (x) in accordance with the instructions given on the back of the answer sheet.

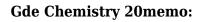
Universal gas constant $R = 8.314 \text{ J K}^{-1} \text{ mol}^{-1}$ Avogadro constant $N_a = 6.022 \times 10^{23} \text{ mol}^{-1}$ Planck's constant $h = 6.626 \times 10^{-34} \text{ J s}$ Velocity of light $c = 3 \times 10^8 \text{ m s}^{-1}$

- I. Which of the following electronic transitions belongs to the visible region of the line spectrum of atomic hydrogen? (n = Principal quantum number)
 - (1) $n=5 \rightarrow n=3$ (2) $n=4 \rightarrow n=2$
 - (3) $n = 1 \rightarrow n = 2$
 - (4) $n=3 \rightarrow n=1$
 - (5) $n=2 \rightarrow n=1$
- 2. Select the incorrect statement.
 - (1) Pauli exclusion principle excludes the possibility of the presence of a third electron in an orbital.
 - (2) In a potassium atom, the number of electrons having quantum numbers n (principal quantum number) = 3 and m, (magnetic quantum number) = 0 is 4.
 - (3) The effective nuclear charge felt by a valence electron in nitrogen (N), is greater than the effective nuclear charge felt by a valence electron in carbon (C).
 - (4) Among the ions Na*, Mg^{2*}, K* and Ca^{2*}, the two ions closest to each other in size are K* and Mg^{2*}.
 - (5) Electron gain energy of carbon is negative.
- The increasing order of the second ionization energy (X*(g) → X²*(g) + e) of Be, B and O is,
 - (I) Be < B < O
 - (2) Be < O < B (3) B < O < Be
 - (4) B < Be < O
 - (5) O < Be < B
- 4. The shapes of F₄ClO, FClO, and FClO₄ are respectively,
 - (1) tetrahedral, trigonal planar and see-saw.
 - (2) square planar, trigonal planar and tetrahedral.
 - (3) see-saw, trigonal pyramidal and square planar.
 - (4) tetrahedral, trigonal pyramidal and see-saw.
 - (5) see-saw, trigonal pyramidal and tetrahedral.

[See page two

Gde Chemistry 20memo

IM Harris



This book delves into Gde Chemistry 20memo. Gde Chemistry 20memo is an essential topic that needs to be grasped by everyone, from students and scholars to the general public. This book will furnish comprehensive and in-depth insights into Gde Chemistry 20memo, encompassing both the fundamentals and more intricate discussions.

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 - Chapter 2: Essential Elements of Gde Chemistry 20memo
 - o Chapter 3: Gde Chemistry 20memo in Everyday Life
 - Chapter 4: Gde Chemistry 20memo in Specific Contexts
 - ∘ Chapter 5: Conclusion
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- 6. In chapter 5, this book will draw a conclusion about Gde Chemistry 20memo. The final chapter will summarize the key points that have been discussed throughout the book.
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