



Periodic Table of the Elements

1	2											3	4	5	6	7	8	9	10	11	12	13	14	15	16	17	18
H	He											Li	Be	B	C	N	O	F	Ne	Na	Mg	Al	Si	P	S	Cl	Ar
K	Ca											Sc	Ti	V	Cr	Mn	Fe	Co	Ni	Cu	Zn	Ga	Ge	As	Se	Br	Kr
Rb	Sr											Y	Zr	Nb	Mo	Tc	Ru	Rh	Pd	Ag	Cd	In	Sn	Sb	Te	I	Xe
Cs	Ba											La	Hf	Ta	W	Re	Os	Ir	Pt	Au	Hg	Tl	Pb	Bi	Po	At	Rn
Fr	Ra											Ac	Th	Pa	U	Np	Pu	A	Am	Cm	Bk	Cf	Es	Fm	M	Lv	Uu

Legend:

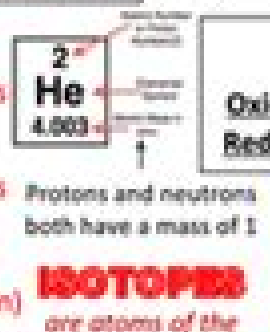
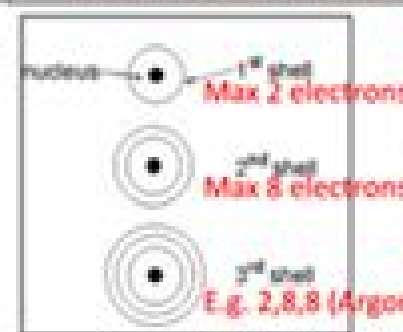
- alkali metals
- alkali earth metals
- transition metals
- post metals
- metalloids
- non metals
- very reactive metals

An atom:
 Number of Protons = Number of Electrons
 Mass number = Protons + Neutrons

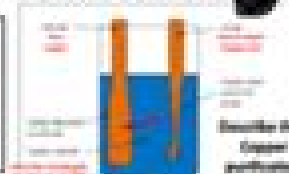
C2 Chemistry

Periodic table arranges elements by
ATOMIC NUMBER (proton number/the small one)

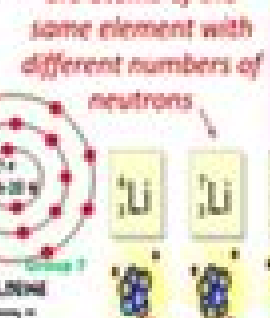
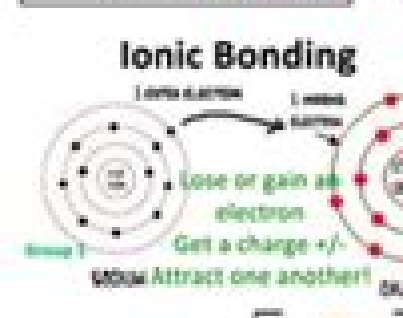
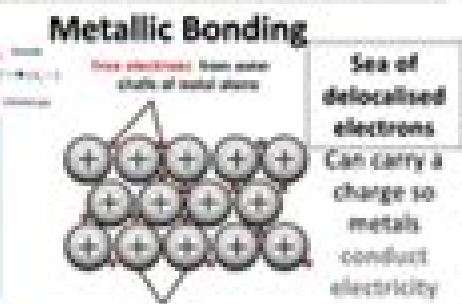
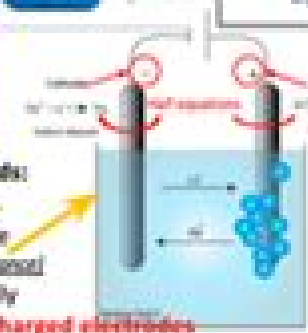
The number of outer shell electrons match the group the element is found in.
 E.g. Lithium 2,1 is a group 1 element.



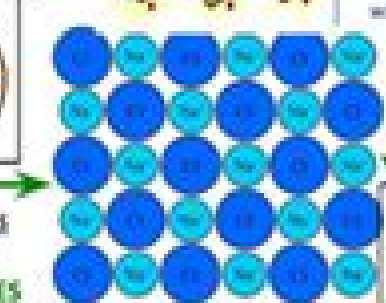
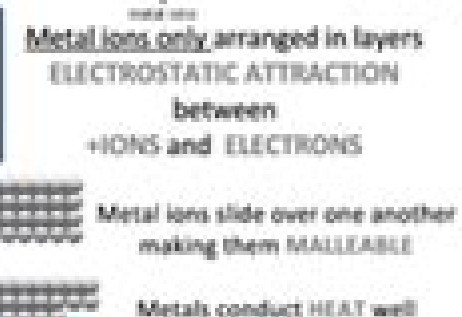
OIL RIG
 Oxidation is loss (of electrons)
 Reduction is gain (of electrons)



Electrolysis:
 Splitting up a substance using electricity



When there are 2 possible products
 The less reactive forms at the electrode



As a solid: High melting and Boiling Point
 Conducts electricity when melted or dissolved: ions move freely and carry a charge

NANO-TECHNOLOGY
 Means REALLY REALLY REALLY small!!!!

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Brendan G. Carr



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