

What's New? A Fun! Simulation about Molecular Geometry and Polarity.

A study of electronegativity, bond polarity, and molecular polarity (part 1)

Introduction:

In this atomic-level simulation, you will investigate how atoms' electronegativity value affects the bond's behavior. When two atoms bond, a pair of electrons is shared between atoms. Electronegativity is a measure of a single atom's ability to hold electrons, shared or that bond. In this lab you will work diligently, at your own pace, to answer a number of questions.

If a ligand, from what you've already learned about the protons and electrons in an atom, what kind of gases are atoms to have a high electronegativity values? High number of valence electrons, mean an atom also small atoms with low shielding.

Why might an atom have a low electronegativity value?

atoms electron cloud shell (L.E.E.) large atoms with lots of shielding.

Procedures:

- Turn on (click) all view options: ORBIT and spin the molecule with your mouse.
- Play your time and investigate how the binary compound's bond behaves when the atom's electronegativity and oxidation are changed. Do not rush through this step.

Describe the polarity of the bond formed between two atoms with similar, low electronegativity.

____nonpolar, no molecular dipole,____

Describe the bond formed between two atoms with similar, high electronegativity.

____nonpolar, no molecular dipole,____

Describe the bond formed between two atoms with very different electronegativity.

____polar, has molecular dipole,____

Describe (in your own words) what is meant by partial charges, δ^- and δ^+ , ____positive, less or negative ions,____

δ^- represent to ____partial negative,____

δ^+ represent to ____partial positive,____

What happens when the electric field is applied to a very polar molecule? ____molecule align itself with the field,____

Why do you think this is? ____opposites attract and likes repel,____

What is electron density? ____electron cloud is more attracted to one atom or side of a molecule,____

How does the density around a partial positive compare to the density of a partial negative? ____

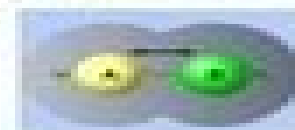
____partial positive density < partial negative density,____

What would bring about a higher electron density around an atom? ____higher electronegativity value,____

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Name	
<input type="checkbox"/> All of them, even	
<input type="checkbox"/> More Chlorine	
<input type="checkbox"/> More Oxygen	
Surface	
<input type="radio"/> None	
<input type="checkbox"/> Electronegativity	
<input type="checkbox"/> Partial Charge	
Electric Field	
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