

“What Do you Think?” A PhET Simulation about Molecular Geometry and Polarity.

A study of electronegativity, bond polarity, and molecular polarity (part 1)

Introduction:

In this atomic-level simulation, you will investigate how atoms' electronegativity value affects the bond's behavior. When two atoms bond, a pair of electrons is shared between atoms. Electronegativity is a measure of a single atom's ability to hold electrons, shared in that bond. In this lab you will work diligently, at your own pace, to answer a number of questions.

If a bond, from what you've already learned about the protons and electrons in an atom, what kind of species an atom is likely to have a high electronegativity value? A high number of valence electrons, mean an atom is also small atoms with low shielding.

Why might an atom have a low electronegativity value?

atoms with more than one shell (3, 4, 5) large atoms with lots of shielding

Procedures:

- Turn on (check) all view options: ORBIT and spin the molecule with your mouse.
- Play your time and investigate how the binary compound's bond behaves when the atom's electronegativity and oxidation are changed. Do not rush through this step.

Describe the polarity of the bond formed between two atoms with similar, low electronegativity.

____nonpolar, no molecular dipole____

Describe the bond formed between two atoms with similar, high electronegativity.

____nonpolar, no molecular dipole____

Describe the bond formed between two atoms with very different electronegativity.

____polar, has molecular dipole____

Describe (in your own words) what is meant by partial charges, δ^- and δ^+ . ____positive, less or negative ions____

δ^- represents: ____partial negative____

δ^+ represents: ____partial positive____

What happens when the electric field is applied to a very polar molecule? ____molecule aligns itself with the field____

Why do you think this is? ____opposites attract and like repel____

What is electron density? ____electron cloud is more attracted to one atom or side of a molecule____

How does the density around a partial positive compare to the density of a partial negative? ____

____partial positive density < partial negative density____

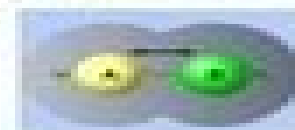
What would bring about a higher electron density around an atom? ____higher electronegativity value____

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Introduction to Polarity

Name	
<input type="checkbox"/> Show Bond Order	
<input checked="" type="checkbox"/> Show Charge	
<input checked="" type="checkbox"/> Show Dipole	
Rotation	
<input type="radio"/> None	
<input type="radio"/> Rotate Around Bond	
<input checked="" type="radio"/> Rotate Around Center	
Electron Density	
<input checked="" type="checkbox"/> On	



Molecule Polarity Phet Lab Teachers Answer Key

LM Reder

A decorative graphic consisting of a light blue horizontal bar with a rounded right end, and a red circular gradient shape partially visible behind it.

Molecule Polarity Phet Lab Teachers Answer Key:

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