

Date: \_\_\_\_\_ Block: \_\_\_\_\_ Name: \_\_\_\_\_

**Chemistry 11 - MOLE RATIOS (using the three step method)**

Magnesium metal reacts with hydrochloric acid to produce magnesium chloride and hydrogen gas. The chemical equation is given as the following:



If 4 moles of HCl is used, how many moles of  $\text{MgCl}_2$  will be formed in the reaction?

Step (1): What is the mole ratio for HCl to  $\text{MgCl}_2$ ?



Step (2): Write two conversion factors for this mole ratio.



Step (3): Calculate the number of moles of  $\text{MgCl}_2$  formed.



*Now try these example questions on a separate sheet of paper using the 3-step method!*



- a) If 0.65 moles of  $\text{CaSi}_2$  is used, how many moles of Sb will be formed in the reaction?

Answer: 0.43 mol Sb

- b) If 1.47 moles of Si is formed in the reaction, how many moles of  $\text{SbCl}_3$  was used at the start of the reaction?

Answer: 0.49 mol  $\text{SbCl}_3$

2. In a Chemistry 11 experiment, John mixes 0.010 moles of calcium chloride together with silver nitrate in a beaker. This results in the formation of silver chloride and calcium nitrate. How many moles of silver chloride are formed in the reaction?

Answer: 0.020 mol AgCl

# Mole Ratio Answers

**M Woodhall**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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## **Table of Contents Mole Ratio Answers**

1. Understanding the eBook Mole Ratio Answers
  - The Rise of Digital Reading Mole Ratio Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Mole Ratio Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Mole Ratio Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Mole Ratio Answers
  - Personalized Recommendations
  - Mole Ratio Answers User Reviews and Ratings
  - Mole Ratio Answers and Bestseller Lists
5. Accessing Mole Ratio Answers Free and Paid eBooks
  - Mole Ratio Answers Public Domain eBooks
  - Mole Ratio Answers eBook Subscription Services
  - Mole Ratio Answers Budget-Friendly Options
6. Navigating Mole Ratio Answers eBook Formats

- ePub, PDF, MOBI, and More
- Mole Ratio Answers Compatibility with Devices
- Mole Ratio Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Mole Ratio Answers
  - Highlighting and Note-Taking Mole Ratio Answers
  - Interactive Elements Mole Ratio Answers
- 8. Staying Engaged with Mole Ratio Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Mole Ratio Answers
- 9. Balancing eBooks and Physical Books Mole Ratio Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Mole Ratio Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Mole Ratio Answers
  - Setting Reading Goals Mole Ratio Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Mole Ratio Answers
  - Fact-Checking eBook Content of Mole Ratio Answers
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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