

## Moles and Chemical Formulas

11

## A. Finding the Simplest Formula

1. Mass of empty crucible + cover
2. Initial appearance of the magnesium
3. Mass of crucible + cover + magnesium
4. Mass of crucible + cover + oxide product

	<u>22.842</u>	g
	<u>23.483</u>	
	<u>23.073</u>	g
	<u>29.317</u>	g

## Calculations

5. Mass of magnesium
6. Mass of magnesium compound
7. Mass of oxygen in the product
8. Moles of Mg

	<u>0.2305</u>	g
	<u>6.475</u>	g
	<u>6.245</u>	g
	<u>.0095</u>	g

(Show calculations.)

$$2305 \text{ g/mg} \times \frac{1 \text{ mole mg}}{24.31 \text{ g/mg}} = .009481$$

9. Moles of O

(Show calculations.)

$$6.245 \text{ g} \times \frac{1 \text{ mole}}{16.00} = .3903$$

10. Which number of moles (Mg or O) is smaller Mole of Mg  
moles of Mg (rounded)

# Mole Lab Chemistry I Acc Answers

**BM King**



## **Mole Lab Chemistry I Acc Answers:**

## Decoding **Mole Lab Chemistry I Acc Answers**: Revealing the Captivating Potential of Verbal Expression

In an era characterized by interconnectedness and an insatiable thirst for knowledge, the captivating potential of verbal expression has emerged as a formidable force. Its capability to evoke sentiments, stimulate introspection, and incite profound transformations is genuinely awe-inspiring. Within the pages of "**Mole Lab Chemistry I Acc Answers**," a mesmerizing literary creation penned with a celebrated wordsmith, readers embark on an enlightening odyssey, unraveling the intricate significance of language and its enduring affect our lives. In this appraisal, we shall explore the book is central themes, evaluate its distinctive writing style, and gauge its pervasive influence on the hearts and minds of its readership.

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