

## Unit: Stoichiometry Answer key "Mole-Mole Calculations" - WS #1

Directions: For each of the following problems, show your balanced chemical equation and all of your calculations.

 How many moles of potassium oxide (K<sub>2</sub>O) will be formed when 1.52 moles of potassium reacts with oxygen according to the following reaction:

4 K + O<sub>2</sub> → 2 K<sub>2</sub>O  
1.52 mol K x 
$$\frac{2 \mod K_2 O}{4 \mod K}$$
 = 0.760 mol K<sub>2</sub>O

How many moles of copper are needed to react with sulfur to produce 0.25 moles of copper(I) sulfide according to the following reaction:

2 Cu + S 
$$\rightarrow$$
 Cu<sub>2</sub>S  
0.25 moi Cu<sub>2</sub>S x  $\frac{2 \ moi \ Cu}{1 \ moi \ Cu_2S}$  = 0.50 moi Cu

How many moles of sodium are needed to react with sulfuric acid to produce 3.75 moles of sodium sulfate according to the following (unbalanced) reaction:

\_\_\_2\_ Na + \_\_\_ H<sub>2</sub>(SO<sub>4</sub>) 
$$\rightarrow$$
 \_\_\_\_ Na<sub>2</sub>(SO<sub>4</sub>) + \_\_\_\_ H<sub>2</sub>  
3.75 mol Na<sub>2</sub>SO<sub>4</sub> x  $\frac{2 \text{ mol Na}}{1 \text{ mol Na}_2SO_4}$  = 7.50 mol Na

How many moles of calcium oxide will be produced when 1.6 moles of Iron(III)oxide reacts with calcium phosphate according to the following (unbalanced) reaction:

Fe<sub>2</sub>O<sub>3</sub> + \_\_\_\_ Ca<sub>3</sub>(PO<sub>4</sub>)<sub>2</sub> 
$$\rightarrow$$
 \_\_\_2\_ Fe(PO<sub>4</sub>) + \_\_\_3\_ CaO  
1.6 mol Fe<sub>2</sub>O<sub>3</sub> x  $\frac{3 mol CaO}{1 mol Fe2O3}$  = 4.8 mol Ca

4. How many moles of oxygen are produced when 5.12 moles of magnesium chlorate decomposes into magnesium chloride and oxygen?

Word equation: magnesium chlorate → magnesium chloride + oxygen

Formula equation: 
$$Mg(ClO_3)_2 \rightarrow MgCl_2 + 3O_2$$
  
(must be balanced!)
$$5.12 \mod Mg(ClO_3)2 \times \frac{3 \mod O_2}{1 \mod Mg(ClO_3)2} = 15.4 \mod O_2$$

How many moles of Chromium(III) nitrate are produced when Chromium reacts with 0.85 moles of Lead(IV) nitrate to produce chromium(III) nitrate and lead?

Word equation: chromium + lead (IV) nitrate → chromium (III) nitrate + lead

Formula equation: 
$$4\text{Cr} + 3\text{Pb}(\text{NO}_3)_4 \Rightarrow 4\text{Cr}(\text{NO}_3)_3 + 3\text{Pb}$$
 (must be balanced!)

$$0.85 \ mol \ Pb(\text{NO}_3)_4 \times \frac{4 \ mol \ Cr(\text{NO}_3)_3}{3 \ mol \ Pb(\text{NO}_3)_4} = 1.1 \ mol \ Cr(\text{NO}_3)_3$$

# **Mole And Stoichiometry Practice Answer Key**

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