

Moz

Molarity, Molality, & Solutions Practice

Molarity $M = \frac{\text{moles}}{L}$

- Find the M of 1 mole of salt in 2.5 L. of water.

$$\frac{1 \text{ mole}}{2.5 L} = \boxed{.4 M}$$

- Find the M of 2.3 moles of salt in 1.14L. of water.

$$\frac{2.3 \text{ moles}}{1.14} = \boxed{2 M}$$

- Find the M of 1.9 moles of sodium acetate in 12.4L. of water.

$$\frac{1.9}{12.4} = \boxed{.15 M}$$

- How many moles of potassium carbonate are needed to make 500 mL. of a 1.25 M solution?

$$1.25 M = \frac{\text{moles}}{.5 L} \quad (.5)(1.25) = \text{moles} \quad \boxed{0.625 \text{ moles}}$$

- How many liters of 2.5 M solution can be made using 1.8 moles of lithium bromide?

$$2.5 M = \frac{1.8 \text{ moles}}{V} \quad \frac{(2.5)(V)}{2.5} = \frac{1.8}{2.5} \quad \boxed{0.72 L}$$

- What is the concentration of a 450 mL. solution that contains 200 grams of iron (II) chloride?

$$\frac{200 \text{ g FeCl}_2}{112.15 \text{ g}} \left| \frac{1 \text{ mole FeCl}_2}{112.15 \text{ g}} \right| = 1.77 \text{ moles} \quad \frac{1.77}{.45 L} = \boxed{3.93 M}$$

- How many grams of ammonium sulfate are needed to make 0.75L. of solution at a concentration of 4.0 M?

$$4 M = \frac{\text{moles}}{0.75} = 3 \text{ moles} \quad \frac{3 \text{ moles } (NH_4)_2SO_4}{132.17 \text{ g}} \left| \frac{1 \text{ mole } (NH_4)_2SO_4}{132.17 \text{ g}} \right| = \boxed{396.5 \text{ g}}$$

Molality

- Calculate the molality of a solution containing 2.89g of NaCl dissolved in 0.159 L. of water.

$$\frac{2.89 \text{ g NaCl}}{58.44 \text{ g}} \left| \frac{1 \text{ mole}}{58.44 \text{ g}} \right| = 0.049 \text{ mol} \quad \frac{0.049}{0.159} = \boxed{.31 m}$$

- Calculate the molality of a solution containing 1.80 mol KCl in 16.0 mol of H₂O

$$\frac{16 \text{ mol H}_2\text{O}}{1 \text{ mol}} \left| \frac{18 \text{ g}}{1 \text{ mol}} \right| = 288.32 \text{ g} = \text{mol} \quad \frac{1.8 \text{ mol}}{288 \text{ g}} = \boxed{6.25 m}$$

Molarity Practice Unit Lesson Key

David Halliday, Robert Resnick, Jearl Walker



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