

# Chemistry: Molarity



Directions: Using these formulas, solve the following problems:

molarity = mol solute / L solution

1 amu = 1 gram / mol

$M_1V_1 = M_2V_2$

Problem 1: Calculate the molarity of a 2.21 L solution made with 31.7 g of calcium chloride. Answer: \_\_\_\_\_

Problem 2: How many grams of NaCl would be required to produce a 6.1 M solution with a volume of 0.64 L? Answer: \_\_\_\_\_

Problem 3: How many grams of NaOH would be required to produce a 5.36 M solution with a volume of 1.37 L? Answer: \_\_\_\_\_

Problem 4: How many grams of KCl would you need to weigh out to prepare 0.416 L of a 3.37 M KCl solution? Answer: \_\_\_\_\_

Problem 5: Calculate the molarity of a 6.68 mL solution of 6.89 g NaCl. Answer: \_\_\_\_\_

Problem 6: Calculate the volume in mL if you have 10.9 g of  $\text{Na}_2\text{CO}_3$  from a 0.112 M solution. Answer: \_\_\_\_\_

Problem 7: Calculate the volume in mL if you have 28.1 g LiOH from a 7.11 M solution. Answer: \_\_\_\_\_

Problem 8: Calculate the final concentration if water is added to 0.789 L of a 3.68 M HCl solution to give a volume of 0.387 L. Answer: \_\_\_\_\_

Problem 9: Calculate the final concentration if a 8.63 mL of a 1.12 M HCl solution is diluted with water to 2.54 mL. Answer: \_\_\_\_\_

Problem 10: Calculate the final concentration if water is added to 0.235 L of a 6.11 M KBr solution to give a volume of 5.79 L. Answer: \_\_\_\_\_

# Molarity Worksheet Answers Chemistry If8766

**Chao Zhang**



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## **Table of Contents Molarity Worksheet Answers Chemistry If8766**

1. Understanding the eBook Molarity Worksheet Answers Chemistry If8766
  - The Rise of Digital Reading Molarity Worksheet Answers Chemistry If8766
  - Advantages of eBooks Over Traditional Books
2. Identifying Molarity Worksheet Answers Chemistry If8766
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Molarity Worksheet Answers Chemistry If8766
  - User-Friendly Interface
4. Exploring eBook Recommendations from Molarity Worksheet Answers Chemistry If8766
  - Personalized Recommendations
  - Molarity Worksheet Answers Chemistry If8766 User Reviews and Ratings
  - Molarity Worksheet Answers Chemistry If8766 and Bestseller Lists
5. Accessing Molarity Worksheet Answers Chemistry If8766 Free and Paid eBooks

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  - Molarity Worksheet Answers Chemistry If8766 eBook Subscription Services
  - Molarity Worksheet Answers Chemistry If8766 Budget-Friendly Options
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- ePub, PDF, MOBI, and More
  - Molarity Worksheet Answers Chemistry If8766 Compatibility with Devices
  - Molarity Worksheet Answers Chemistry If8766 Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Molarity Worksheet Answers Chemistry If8766
  - Highlighting and Note-Taking Molarity Worksheet Answers Chemistry If8766
  - Interactive Elements Molarity Worksheet Answers Chemistry If8766
8. Staying Engaged with Molarity Worksheet Answers Chemistry If8766
- Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Molarity Worksheet Answers Chemistry If8766
9. Balancing eBooks and Physical Books Molarity Worksheet Answers Chemistry If8766
- Benefits of a Digital Library
  - Creating a Diverse Reading Collection Molarity Worksheet Answers Chemistry If8766
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Molarity Worksheet Answers Chemistry If8766
- Setting Reading Goals Molarity Worksheet Answers Chemistry If8766
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Molarity Worksheet Answers Chemistry If8766
- Fact-Checking eBook Content of Molarity Worksheet Answers Chemistry If8766
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks

#### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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