

Molarity Practice problems

$$M = \frac{\text{mol}}{L}$$



Key

1. How many moles of CaCl_2 would be dissolved in 1.0 L of a 0.1 M solution of CaCl_2 ?

$$\text{mol} = M \times L = 0.1 \frac{\text{mol}}{L} \times 1 L = 0.1 \text{ mol } \text{CaCl}_2$$

2. What is the concentration of 25 mL of solution that has 0.05 moles of solute dissolved in it?

$$M = \frac{\text{mol}}{L} = \frac{0.05 \text{ mol}}{0.025 L} = 2 M$$

3. How many liters of water need to be added to 0.5 moles of CaCl_2 to make a 0.20 M solution of CaCl_2 ?

$$L = \frac{\text{mol}}{M} = \frac{0.5 \text{ mol}}{0.20 M} = 2.5 L$$

4. How many grams of potassium carbonate are needed to make 200 mL of a 2.5 M solution?

$$\text{mol} = M \times L \quad \text{K}_2\text{CO}_3$$

$$\text{mol} = 2.5 M \times 0.200 L = 0.50 \text{ mol}$$

$$\frac{0.50 \text{ mol } \text{K}_2\text{CO}_3}{1 \text{ mol}} \times 138.21 \text{ g} = 69.1 \text{ g } \text{K}_2\text{CO}_3$$

5. How many grams of Lithium Bromide would be needed to make 1.43 L of 4 M Lithium Bromide solution?

$$\text{mol} = M \times L \quad \text{LiBr}$$

$$= 4 M \times 1.43 L = 5.72 \text{ mol } \text{LiBr}$$

$$\frac{5.72 \text{ mol } \text{LiBr}}{1 \text{ mol}} \times 86.84 \text{ g} = 496.7 \text{ g } \text{LiBr}$$

6. What is the concentration of a 450 mL solution that contains 200 grams of iron (II) chloride? FeCl_2

$$M = \frac{\text{mol}}{L}$$

$$\frac{200 \text{ g } \text{FeCl}_2}{126.75 \text{ g}} \times 1 \text{ mol} = 1.578 \text{ mol } \text{FeCl}_2$$

$$\frac{1.578 \text{ mol}}{0.450 L} = 3.5 M$$

7. How many grams of ammonium sulfate are needed to make a 0.25 L solution at a concentration of 6 M?

$$\text{mol} = M \times L \quad (\text{NH}_4)_2\text{SO}_4$$

$$= 6 M \times 0.25 L = 1.5 \text{ mol } (\text{NH}_4)_2\text{SO}_4$$

$$\frac{1.5 \text{ mol } (\text{NH}_4)_2\text{SO}_4}{1 \text{ mol}} \times 132.138 \text{ g} = 198.2 \text{ g } (\text{NH}_4)_2\text{SO}_4$$

8. What is the concentration of a solution that has a volume of 2.5 L and contains 660 grams of calcium phosphate?

$$M = \frac{\text{mol}}{L}$$

$$\frac{660 \text{ g } \text{Ca}_3(\text{PO}_4)_2}{310.18 \text{ g}} \times 1 \text{ mol} = 2.127 \text{ mol}$$

$$\frac{2.127 \text{ mol}}{2.5 L} = 0.85 M$$

9. How many grams of copper (II) fluoride are needed to make 6.7 liters of a 1.2 M solution?

$$\text{mol} = M \times L \quad \text{CuF}_2$$

$$= 1.2 M \times 6.7 L = 8.04 \text{ mol } \text{CuF}_2$$

$$\frac{8.04 \text{ mol } \text{CuF}_2}{1 \text{ mol}} \times 101.55 \text{ g} = 816.5 \text{ g } \text{CuF}_2$$

Molarity Practice Unit 10 Lesson Key

D Keegan

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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