

Molarity Practice

Calculate for the unknown variables. $M = n/v$

1. Explain what molarity/concentration means?
2. What is the molarity of a solution that contains 5.50 moles of calcium chloride in 2.5 liters of solution?
3. What is the molarity of a solution that contains 3.39×10^{-2} moles hydrobromic acid in 1 liter of solution?
4. A beaker contains 250g $C_6H_{12}O_6$ (glucose) in 1,500 mL of solution. What is the molarity?
5. A flask contains 325g grams iron II oxide and has a molarity of 2M. What is the volume of the solution?
6. Calculate the moles if a flask contains 3.25M potassium nitrate in 750 mL of solution.
7. A cup contains 85.31 grams ammonium acetate in 3.27 liters of solution - calculate the molarity.
8. What is the molarity of a solution that contains 10.5 grams magnesium oxide in 250 mL of solution?
9. What mass of potassium phosphate would you use to make 2.5 liter of 1.25 M solution?

Molarity Practice Chemistry Unit 10 Lesson

Liyong Dong



Molarity Practice Chemistry Unit 10 Lesson :

Reviewing **Molarity Practice Chemistry Unit 10 Lesson** : Unlocking the Spellbinding Force of Linguistics

In a fast-paced world fueled by information and interconnectivity, the spellbinding force of linguistics has acquired newfound prominence. Its capacity to evoke emotions, stimulate contemplation, and stimulate metamorphosis is actually astonishing. Within the pages of "**Molarity Practice Chemistry Unit 10 Lesson** ," an enthralling opus penned by a very acclaimed wordsmith, readers attempt an immersive expedition to unravel the intricate significance of language and its indelible imprint on our lives. Throughout this assessment, we shall delve in to the book is central motifs, appraise its distinctive narrative style, and gauge its overarching influence on the minds of its readers.

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