

9) A 250mL solution of 0.5M NaF has 50mL of a 1M BaF₂ solution added to it. What is the concentration of [F⁻] in the new solution? Assume that volume is additive.

From (8)

$$[F^-] = 0.125 \text{ mol} \left| \begin{array}{l} 0.05 \text{ L} \times 1 \text{ M BaF}_2 \\ = 0.05 \text{ mol BaF}_2 \\ = 0.1 \text{ mol F}^- \end{array} \right| \begin{array}{l} \text{Final sol'n} \\ \text{mol [F]}^- = 0.125 \text{ mol} + 0.1 \text{ mol} = 0.225 \text{ mol F}^- \\ \text{volume} = 0.25 \text{ L} + 0.05 \text{ L} \\ = 0.3 \text{ L} \end{array}$$

$$M = \frac{0.225 \text{ mol}}{0.3 \text{ L}} = \boxed{0.75 \text{ M}}$$

10) How many grams of NaHCO₃ (molar mass = 84.01 g/mol) are needed to make a 350mL solution of 0.75M NaHCO₃?

$$n = V \times M$$

$$= 0.35 \text{ L} \times 0.75 \text{ M NaHCO}_3$$

$$= 0.2625 \text{ mol NaHCO}_3$$

$$\text{mass} = 0.2625 \text{ mol NaHCO}_3 \times \frac{84.01 \text{ g NaHCO}_3}{1 \text{ mol NaHCO}_3}$$

$$= \boxed{22.05 \text{ g NaHCO}_3}$$

11) Calculate the amount of distilled water that is needed to dilute a 3L solution of 5M HBr to a concentration of 1.5M HBr?

$$\text{mol} = 3 \text{ L} \times 5 \text{ M}$$

$$= 15 \text{ mol HBr}$$

volume needed: $V = 15 \text{ mol} / 1.5 \text{ M HBr}$

$$= 10 \text{ L}$$

$$10 \text{ L} - 3 \text{ L} = \boxed{7 \text{ L needed to add}}$$

12) What volume of 8M H₂SO₄ is needed to prepare a 100mL solution of 0.5M H₂SO₄?

$$0.1 \text{ L} \times 0.5 \text{ M H}_2\text{SO}_4$$

$$= 0.005 \text{ mol H}_2\text{SO}_4 \text{ needed}$$

$$V_{\text{ol}} = \frac{0.005 \text{ mol H}_2\text{SO}_4}{8 \text{ M H}_2\text{SO}_4}$$

$$= 6.25 \times 10^{-4} \text{ L}$$

$$= \boxed{0.625 \text{ mL of } 8 \text{ M H}_2\text{SO}_4, \text{ the rest is water}}$$

13) The maximum amount of salt that can be contained in water is approximately 26.3% by weight. Convert this into concentration with units of molarity. Info: 1g of water = 1 mL of water.

assume 100g

$$26.3 \text{ g NaCl}$$

$$26.3 \text{ g NaCl} \times \frac{1 \text{ mol NaCl}}{58.44 \text{ g NaCl}} = 0.45 \text{ mol NaCl}$$

$$73.7 \text{ g H}_2\text{O} \times \frac{1 \text{ mol H}_2\text{O}}{18 \text{ g H}_2\text{O}} = 4.1 \text{ mol H}_2\text{O}$$

$$M = \frac{0.45 \text{ mol NaCl}}{0.0737 \text{ L sol'n}} = \boxed{6.1 \text{ M}}$$

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Yan Bai

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web the midnight of your birth 5 carols for upper voi shadows under the midnight sun oct 23 2020 hans laub fell back but he kept his balance the norwegian continued his advance his fist loaded and aimed at its retreating target stop swine hans laub bellowed don t come near me but the norwegian came a step behind laub s