

Molarity Calculations

Calculate the molarities of the following solutions:



- 1) 2.3 moles of sodium chloride in 0.45 liters of solution.

$$M = \frac{2.3 \text{ moles}}{0.45 \text{ L}} = 5.111 \frac{\text{mol}}{\text{L}} = \boxed{5.11 \text{ M}}$$



- 2) 1.2 moles of calcium carbonate in 1.22 liters of solution.

$$M = \frac{1.2 \text{ mol}}{1.22 \text{ L}} = 0.983606557 \frac{\text{mol}}{\text{L}} = \boxed{0.98 \text{ M}}$$



- 3) 0.09 moles of sodium sulfate in 12 mL of solution.

12 mL = 0.012 L

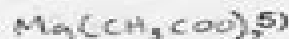
$$M = \frac{0.09 \text{ mol}}{0.012 \text{ L}} = \boxed{7.5 \text{ M}}$$



- 4) 0.75 moles of lithium fluoride in 65 mL of solution.

65 mL = 0.065 L

$$M = \frac{0.75 \text{ mol}}{0.065 \text{ L}} = 11.53846154 \text{ M} = \boxed{11.54 \text{ M}}$$



- 5) 0.8 moles of magnesium acetate in 5 liters of solution.

$$M = \frac{0.8 \text{ mol}}{5 \text{ L}} = \boxed{0.16 \text{ M}}$$



- 6) 120 grams of calcium nitrite in 240 mL of solution.

240 mL = 0.240 L

$$n = 120 \text{ g} \left(\frac{1 \text{ mol}}{132.09 \text{ g}} \right)$$

MW_{Ca(NO₂)₂} = 132.09 g/mol

n = 0.91 mol

$$M = \frac{0.91 \text{ mol}}{0.240 \text{ L}} = \boxed{3.79 \text{ M}}$$



- 7) 98 grams of sodium hydroxide in 2.2 liters of solution.

MW_{NaOH} = 39.997 g/mol

$$n = 98 \text{ g} \left(\frac{1 \text{ mol}}{39.997 \text{ g}} \right) = 2.45 \text{ mol}$$

$$M = \frac{2.45 \text{ mol}}{2.2 \text{ L}} = \boxed{1.11 \text{ M}}$$



- 8) 1.2 grams of hydrochloric acid in 25 mL of solution.

25 mL = 0.025 L

MW_{HCl} = 36.461 g/mol

$$n = (1.2 \text{ g}) \left(\frac{1 \text{ mol}}{36.461 \text{ g}} \right) = 0.033 \text{ mol}$$

$$M = \frac{0.033 \text{ mol}}{0.025 \text{ L}} = \boxed{1.32 \text{ M}}$$



- 9) 45 grams of ammonia in 0.75 L of solution.

MW_{NH₃} = 17.031 g/mol

$$n = 45 \text{ g} \left(\frac{1 \text{ mol}}{17.031 \text{ g}} \right) = 2.64 \text{ mol}$$

$$M = \frac{2.64 \text{ mol}}{0.75 \text{ L}} = \boxed{3.52 \text{ M}}$$

Molarity Calculations Key Answers Show Work

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Chemistry: 1,001 Practice Problems For Dummies (+ Free Online Practice) Heather Hattori, Richard H.

Langley, 2014-03-11 Practice makes perfect and helps deepen your understanding of chemistry Every high school requires a course in chemistry and many universities require the course for majors in medicine engineering biology and various other sciences 1001 Chemistry Practice Problems For Dummies provides students of this popular course the chance to practice what they learn in class deepening their understanding of the material and allowing for supplemental explanation of difficult topics 1001 Chemistry Practice Problems For Dummies takes you beyond the instruction and guidance offered in Chemistry For Dummies giving you 1 001 opportunities to practice solving problems from the major topics in chemistry Plus an online component provides you with a collection of chemistry problems presented in multiple choice format to further help you test your skills as you go Gives you a chance to practice and reinforce the skills you learn in chemistry class Helps you refine your understanding of chemistry Practice problems with answer explanations that detail every step of every problem Whether you re studying chemistry at the high school college or graduate level the practice problems in 1001 Chemistry Practice Problems For Dummies range in areas of difficulty and style providing you with the practice help you need to score high at exam time [E3 Chemistry Guided Study Book - 2018 Home Edition \(Answer Key Included\)](#) Effiong Eyo, 2017-12-08

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Home Edition Also available in paperback print *Analytical Chemistry* Bryan M. Ham,Aihui MaHam,2024-02-28

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Chemistry: The Central Science Theodore L. Brown, H. Eugene LeMay Jr., Bruce

E. Bursten, Catherine Murphy, Patrick Woodward, Steven Langford, Dalius Sagatys, Adrian George, 2013-10-04 If you think you know the Brown LeMay Bursten Chemistry text think again In response to market request we have created the third Australian edition of the US bestseller Chemistry The Central Science An extensive revision has taken this text to new heights Triple checked for scientific accuracy and consistency this edition is a more seamless and cohesive product yet retains the clarity innovative pedagogy functional problem solving and visuals of the previous version All artwork and images are now consistent in quality across the entire text And with a more traditional and logical organisation of the Organic Chemistry content this comprehensive text is the source of all the information and practice problems students are likely to need for conceptual understanding development of problem solving skills reference and test preparation Linne & Ringsrud's Clinical Laboratory Science - E-Book Mary Louise Turgeon, 2015-02-10 Using a discipline by discipline approach Linne perforated pages make it easy for students to remove for use in the lab and for assignment of review questions as homework NEW Instrumentation updates show new technology being used in the lab NEW Additional key terms in each chapter cover need to know terminology NEW Additional tables and figures in each chapter clarify clinical lab science concepts *Clinical Laboratory Science - E-Book* Mary Louise Turgeon, 2022-09-14 Selected for Doody's Core Titles 2024 in Laboratory Technology Using a discipline by discipline approach Turgeon's Clinical Laboratory Science Concepts Procedures and Clinical Applications 9th Edition provides a fundamental overview of the concepts procedures and clinical applications essential for working in a clinical laboratory and performing routine clinical lab tests Coverage includes basic laboratory techniques and key topics such as safety phlebotomy quality assessment automation and point of care testing as well as discussion of clinical laboratory specialties Clear straightforward instructions simplify laboratory procedures and are guided by the latest practices and CLSI Clinical and Laboratory Standards Institute standards Written by well known CLS educator Mary Louise Turgeon this edition offers essential guidance and recommendations for today's laboratory testing methods and clinical applications Broad scope of coverage makes this text an ideal companion for clinical laboratory science programs at various levels including CLS MT CLT MLT medical laboratory assistant and medical assisting and reflects the taxonomy levels of the CLS MT and CLT MLT exams Detailed procedure guides and procedure worksheets on Evolve and in the ebook familiarize you with the exact steps performed in the lab Vivid full color illustrations depict concepts and applicable images that can be seen under the microscope An extensive number of certification style multiple choice review questions are organized and coordinated under major topical headings at the end of each chapter to help you assess your understanding and identify areas requiring additional study Case studies include critical thinking group discussion questions providing the opportunity to apply content to real life scenarios The newest Entry Level Curriculum Updates for workforce entry published by the American Society for Clinical Laboratory Science ASCLS and the American Society for Clinical Pathology ASCP Board of Certification Exam Content Outlines serve as content reference sources Convenient glossary makes it easy to look up

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WileyPLUS sold separately from text **Thermodynamics of Solutions** Eli Ruckenstein, Ivan L. Shulgin, 2009-06-17 This

book consists of a number of papers regarding the thermodynamics and structure of multicomponent systems that we have published during the last decade. Even though they involve different topics and different systems, they have something in common which can be considered as the signature of the present book. First, these papers are concerned with difficult or very nonideal systems, i.e. systems with very strong interactions, e.g. hydrogen bonding between components or systems with large differences in the partial molar volumes of the components, e.g. the aqueous solutions of proteins or systems that are far from normal conditions, e.g. critical or near critical mixtures. Second, the conventional thermodynamic methods are not sufficient for the accurate treatment of these mixtures. Last but not least, these systems are of interest for the pharmaceutical, biomedical, and related industries. In order to meet the thermodynamic challenges involved in these complex mixtures, we employed a variety of traditional methods but also new methods such as the fluctuation theory of Kirkwood and Buff and ab initio quantum mechanical techniques. The Kirkwood-Buff (KB) theory is a rigorous formalism which is free of any of the approximations usually used in the thermodynamic treatment of multicomponent systems. This theory appears to be very fruitful when applied to the above-mentioned difficult systems.

Methods in Molecular Biophysics Nathan R. Zaccai, Igor N. Serdyuk, Joseph Zaccai, 2017-05-18

Current techniques for studying biological macromolecules and their interactions are based on the application of physical methods ranging from classical thermodynamics to more recently developed techniques for the detection and manipulation of single molecules. Reflecting the advances made in biophysics research over the past decade and now including a new section on medical imaging, this new edition describes the physical methods used in modern biology. All key techniques are covered, including mass spectrometry, hydrodynamics, microscopy, and imaging, diffraction, and spectroscopy, electron microscopy, molecular dynamics simulations, and nuclear magnetic resonance. Each method is explained in detail using examples of real-world applications. Short asides are provided throughout to ensure that explanations are accessible to life scientists, physicists, and those with medical backgrounds. The book remains an unparalleled and comprehensive resource for graduate students of biophysics and medical physics in science and medical schools as well as for research scientists looking for an introduction to techniques from across this interdisciplinary field.

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