

MOLARITY BY DILUTION

79 + 52 = 131 Name _____

Acids are usually acquired from chemical supply houses in concentrated form. These acids are diluted to the desired concentration by adding water. Since moles of acid before dilution = moles of acid after dilution, and moles of acid = $M \times V$ then,
 $M_1 \times V_1 = M_2 \times V_2$ Solve the following problems.

1. How much concentrated 18 M sulfuric acid is needed to prepare 250 mL of a 6.0 M solution?

$$V_1 = \frac{M_2 V_2}{M_1} = \frac{(6.0 M)(250 mL)}{18 M}$$

83 mL

2. How much concentrated 12 M hydrochloric acid is needed to prepare 100 mL of a 2.0 M solution?

$$V_1 = \frac{M_2 V_2}{M_1} = \frac{(2.0 M)(100 mL)}{12 M}$$

17 mL

3. To what volume should 25 mL of 15 M nitric acid be diluted to prepare a 3.0 M solution?

$$V_2 = \frac{M_1 V_1}{M_2} = \frac{(15 M)(25 mL)}{3.0 M}$$

125 mL

4. To how much water should 50 mL of 12 M hydrochloric acid be added to produce a 4.0 M solution?

$$V_2 = \frac{M_1 V_1}{M_2} = \frac{(12 M)(50 mL)}{4.0 M} = 150 mL - 50 mL$$

100 mL H₂O

5. To how much water should 100 mL of 18 M sulfuric acid be added to prepare a 1.5 M solution?

$$V_2 = \frac{M_1 V_1}{M_2} = \frac{(18 M)(100 mL)}{1.5 M} = 1200 mL - 100 mL = 1100 mL$$

1.1 L

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RJ Shavelson



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Table of Contents Molarity By Dilution Answer Key Page 69

1. Understanding the eBook Molarity By Dilution Answer Key Page 69
 - The Rise of Digital Reading Molarity By Dilution Answer Key Page 69
 - Advantages of eBooks Over Traditional Books
2. Identifying Molarity By Dilution Answer Key Page 69
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Molarity By Dilution Answer Key Page 69
 - User-Friendly Interface
4. Exploring eBook Recommendations from Molarity By Dilution Answer Key Page 69
 - Personalized Recommendations
 - Molarity By Dilution Answer Key Page 69 User Reviews and Ratings
 - Molarity By Dilution Answer Key Page 69 and Bestseller Lists
5. Accessing Molarity By Dilution Answer Key Page 69 Free and Paid eBooks
 - Molarity By Dilution Answer Key Page 69 Public Domain eBooks
 - Molarity By Dilution Answer Key Page 69 eBook Subscription Services
 - Molarity By Dilution Answer Key Page 69 Budget-Friendly Options

6. Navigating Molarity By Dilution Answer Key Page 69 eBook Formats
 - ePub, PDF, MOBI, and More
 - Molarity By Dilution Answer Key Page 69 Compatibility with Devices
 - Molarity By Dilution Answer Key Page 69 Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Molarity By Dilution Answer Key Page 69
 - Highlighting and Note-Taking Molarity By Dilution Answer Key Page 69
 - Interactive Elements Molarity By Dilution Answer Key Page 69
8. Staying Engaged with Molarity By Dilution Answer Key Page 69
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Molarity By Dilution Answer Key Page 69
9. Balancing eBooks and Physical Books Molarity By Dilution Answer Key Page 69
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Molarity By Dilution Answer Key Page 69
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Molarity By Dilution Answer Key Page 69
 - Setting Reading Goals Molarity By Dilution Answer Key Page 69
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Molarity By Dilution Answer Key Page 69
 - Fact-Checking eBook Content of Molarity By Dilution Answer Key Page 69
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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