

## Molarity Practice Worksheet

Find the molarity (concentration) of the following solutions: Molarity = mole/Liters  
Volume must be in liters! 1 liter = 1000 ml

- 1) 2.5 moles of sodium chloride is dissolved to make 0.050 liters of solution.
- 2) 3.0 moles of silver chloride is dissolved in enough water to make a 2.0 liter solution
- 3) 0.60 moles of potassium nitrate is dissolved in 300 milliliters of solution
- 4.) 4.2 moles of lithium fluoride are dissolved in 200 milliliters of solution.
- 5.) 18 grams of hydrochloric acid, HCl is dissolved to make 0.500 liters of solution.
- 6) 33 grams of Beryllium hydride,  $\text{BeH}_2$ , are dissolved in 6.0 Liters of solution.
- 7) 734 grams of lithium sulfate,  $\text{Li}_2\text{SO}_4$  are dissolved to make 2500 mL of solution.
- 8)  $9.0 \times 10^{23}$  molecules of  $\text{CCl}_4$  dissolved to make 500 mL of solution.
- 6) I have two solutions. In the first solution, 1.0 moles of sodium chloride is dissolved to make 1.0 liters of solution. In the second one, 1.0 moles of

# Molarity And Molality Practice Worksheet Answers

**Padhraic Smyth**



## **Molarity And Molality Practice Worksheet Answers:**

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