

Dilutions and Molarity

Dilution is the process used to make a new solution that is **less concentrated** than the original solution by adding more **solvent**.

M_1 = Initial Molarity

V_1 = Initial Volume

$$M_1 V_1 = M_2 V_2$$

M_2 = Final Molarity

V_2 = Final Volume

Examples Questions:

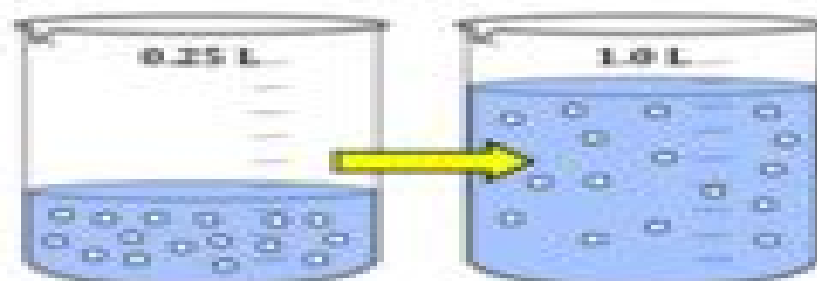
- 1) A scientist obtains 0.25 liters of concentrated 5.0 M NaCl solution. The scientist then **dilutes** the solution by adding water until the solution has a volume of 1.0 liters. What is the **molarity** of the solution?

$$M_1 V_1 = M_2 V_2$$

$$(5.0) (0.25) = M_2 (1.0)$$

$$(1.25) = M_2 (1.0)$$

$$M_2 = 1.25 \text{ M or mol/L}$$



- 2) A concentrated solution of 10.0 M hydrochloric acid solution must be **diluted** to make it weaker. If a scientist starts with 1.5 liters of the original solution and adds 2.5 liters of water, what will be the final **molarity** of the solution?

$$M_1 V_1 = M_2 V_2$$

$$(10.0) (1.5) = M_2 (4.0)$$

$$(15.0) = M_2 (4.0)$$

$$M_2 = 3.75 \text{ M or mol/L}$$

- 3) In a laboratory, a scientist takes 50 mL of 5.0 M NaOH solution and adds water until the solution has a molarity of 2.0 M. What is the **volume** of the new solution?

$$M_1 V_1 = M_2 V_2$$

$$(5.0) (50) = (2.0) V_2$$

$$(250.0) = (2.0) V_2$$

$$V_2 = 125.0 \text{ mL}$$

Molarity And Dilution Worksheet Answers

Raffaella Di Napoli



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Table of Contents Molarity And Dilution Worksheet Answers

1. Understanding the eBook Molarity And Dilution Worksheet Answers
 - The Rise of Digital Reading Molarity And Dilution Worksheet Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Molarity And Dilution Worksheet Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Molarity And Dilution Worksheet Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Molarity And Dilution Worksheet Answers
 - Personalized Recommendations
 - Molarity And Dilution Worksheet Answers User Reviews and Ratings
 - Molarity And Dilution Worksheet Answers and Bestseller Lists
5. Accessing Molarity And Dilution Worksheet Answers Free and Paid eBooks
 - Molarity And Dilution Worksheet Answers Public Domain eBooks
 - Molarity And Dilution Worksheet Answers eBook Subscription Services
 - Molarity And Dilution Worksheet Answers Budget-Friendly Options
6. Navigating Molarity And Dilution Worksheet Answers eBook Formats

- ePub, PDF, MOBI, and More
 - Molarity And Dilution Worksheet Answers Compatibility with Devices
 - Molarity And Dilution Worksheet Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Molarity And Dilution Worksheet Answers
 - Highlighting and Note-Taking Molarity And Dilution Worksheet Answers
 - Interactive Elements Molarity And Dilution Worksheet Answers
 8. Staying Engaged with Molarity And Dilution Worksheet Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Molarity And Dilution Worksheet Answers
 9. Balancing eBooks and Physical Books Molarity And Dilution Worksheet Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Molarity And Dilution Worksheet Answers
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Molarity And Dilution Worksheet Answers
 - Setting Reading Goals Molarity And Dilution Worksheet Answers
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Molarity And Dilution Worksheet Answers
 - Fact-Checking eBook Content of Molarity And Dilution Worksheet Answers
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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