

**Molar Mass Practice 1B – Answer Key****Objectives:** - Calculate molar mass using the periodic table**Directions:** Find the molar mass of the following atoms, molecules, or formula units. Round masses to the hundredths place. Don't forget your units! You must show your work!

1. Ti  $\boxed{47.87 \text{ g/mol}}$



$$2 \times (12.01) + (8 \times 1.01) + 32.07 = \boxed{76.17 \text{ g/mol}}$$

2. Fe  $\boxed{55.85 \text{ g/mol}}$



$$\text{C: } 2 \times 12.01 = 24.02$$

$$\text{H: } 8 \times 1.01 = 8.08$$

$$\text{O: } 1 \times 16.00 = 16.00$$

$$\boxed{48.10 \text{ g/mol}}$$

3.  $\text{CO}_2$   
 $\text{C: } 1 \times 12.01 = 12.01$   
 $\text{O: } 2 \times 16.00 = 32.00$   
$$\boxed{44.01 \text{ g/mol}}$$



$$3\text{Mg: } 3 \times 24.31 = 72.93$$

$$\text{N: } 4 \times 14.01 = 56.04$$

$$\text{O: } 12 \times 16.00 = 192.00$$

$$\boxed{320.97 \text{ g/mol}}$$

4.  $\text{LiCl}$   
 $\text{Li: } 1 \times 6.94 = 6.94$   
 $\text{Cl: } 1 \times 35.45 = 35.45$   
$$\boxed{42.39 \text{ g/mol}}$$



$$2 \times (1.01) + 32.07 + (4 \times 16.00) = \boxed{98.09 \text{ g/mol}}$$

5.  $\text{O}_2$   
 $\text{O: } 2 \times 16.00 = 32.00$   
$$\boxed{32.00 \text{ g/mol}}$$

13.  $\text{Mn}_2(\text{PO}_4)_3$   
 $\text{Mn: } 2 \times 54.94 = 109.88$   
 $\text{P: } 3 \times 30.97 = 92.91$   
 $\text{O: } 12 \times 16.00 = 192.00$   
$$\boxed{394.79 \text{ g/mol}}$$

6. Iron(III) oxide  $\text{Fe}_2\text{O}_3$   
$$2 \times (55.85) + 3 \times (16.00) = \boxed{159.7 \text{ g/mol}}$$



$$2 \times (1.01) + 16.00 = \boxed{18.02 \text{ g/mol}}$$

7.  $\text{HC}_2\text{H}_3\text{O}_2$   
$$\boxed{75.07 \text{ g/mol}}$$

15.  $2\text{KNO}_3$   
$$\boxed{166.1 \text{ g/mol}}$$

8. Lead(II) hydroxide  $\text{Pb}(\text{OH})_2$   
$$\boxed{241.2 \text{ g/mol}}$$

16. Aluminum oxide  $\text{Al}_2\text{O}_3$   
$$\boxed{101.96 \text{ g/mol}}$$

# Molar Mass Practice Answer Key

**Effiong Eyo**

A red circular graphic with a gradient, appearing as a stylized arrow or a partial circle, located to the right of the name bar.

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### **Table of Contents Molar Mass Practice Answer Key**

1. Understanding the eBook Molar Mass Practice Answer Key
  - The Rise of Digital Reading Molar Mass Practice Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Molar Mass Practice Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Molar Mass Practice Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Molar Mass Practice Answer Key
  - Personalized Recommendations
  - Molar Mass Practice Answer Key User Reviews and Ratings

- Molar Mass Practice Answer Key and Bestseller Lists
- 5. Accessing Molar Mass Practice Answer Key Free and Paid eBooks
  - Molar Mass Practice Answer Key Public Domain eBooks
  - Molar Mass Practice Answer Key eBook Subscription Services
  - Molar Mass Practice Answer Key Budget-Friendly Options
- 6. Navigating Molar Mass Practice Answer Key eBook Formats
  - ePub, PDF, MOBI, and More
  - Molar Mass Practice Answer Key Compatibility with Devices
  - Molar Mass Practice Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Molar Mass Practice Answer Key
  - Highlighting and Note-Taking Molar Mass Practice Answer Key
  - Interactive Elements Molar Mass Practice Answer Key
- 8. Staying Engaged with Molar Mass Practice Answer Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Molar Mass Practice Answer Key
- 9. Balancing eBooks and Physical Books Molar Mass Practice Answer Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Molar Mass Practice Answer Key
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Molar Mass Practice Answer Key
  - Setting Reading Goals Molar Mass Practice Answer Key
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Molar Mass Practice Answer Key
  - Fact-Checking eBook Content of Molar Mass Practice Answer Key
  - Distinguishing Credible Sources

13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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