

Molar Mass of a Volatile Liquid

Goal

Determine the molar mass of an unknown volatile liquid.

Background

One of the properties that help characterize a substance is its molar mass. If the substance in question is a volatile liquid, a common method to determine its molar mass is to use the ideal gas law.

$$PV = nRT$$

Because the liquid is volatile, it can easily be converted to a gas. While the substance is in the gas phase, you can measure its volume, pressure, and temperature. You can then use the ideal gas law to calculate the number of moles, n , of the substance. Finally, you can use the number of moles of the gas to calculate molar mass.

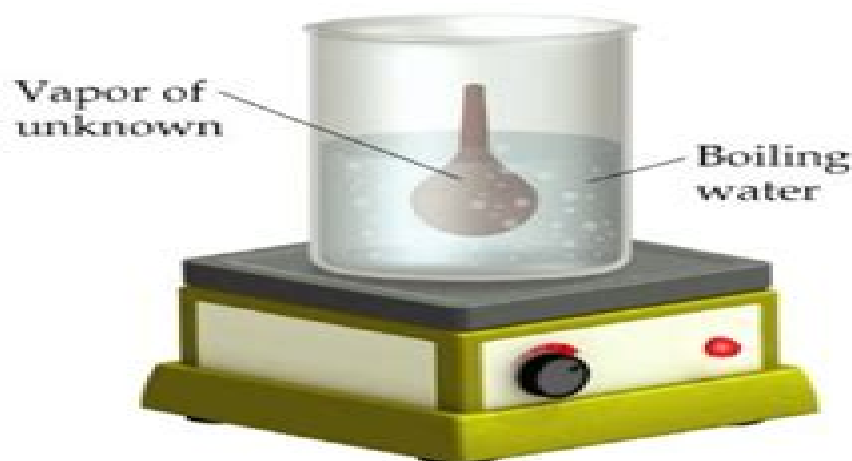


Figure 1

Pre-Laboratory Activity (in lab notebook)

1. Read Ch. 5 in the textbook
2. Identify each variable in the ideal gas Law: $PV = nRT$ (make sure and specify what units each should be in)
3. A volatile liquid is evaporated in a 250.0 mL Dumas at 100.0 °C and 0.98 atm. At these conditions how many moles of gas would fill the Dumas Bulb?
4. Calculate the Molar Mass of the volatile liquid if the mass of gas that filled the Dumas Bulb was 0.736 g.

Reminder: All Measurements should be made with the correct number of significant figures and correct units. All calculations results should have the correct number of significant figures and the correct units.

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SJ Ball



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