

# Chemistry – Unit 6 Notes 1 - Thomson Model of the Atom

J. J. Thomson performed experiments with cathode rays in an attempt to understand electricity – which was still a mystery in the late 1800s. Use the website *A Look Inside the Atom*<sup>1</sup> to find the conclusions that Thomson and other physicists drew regarding the mysterious cathode rays.

Thomson's 1897 Experiments - state the conclusions Thomson drew from each of his famous cathode ray experiments:

1. **First Experiment:** Thomson directed the beam at an electrometer and tried to separate the evidence of charge from the path of the beam. *What connection did Thomson find between charge and the cathode rays? Was the charge positive or negative?*
2. **Second Experiment:** Thomson tried passing the cathode ray through an electric field. *How did cathode ray beam behave when it passed through an electric field? What did he conclude after his second experiment?*
3. **Third Experiment:** Thomson did some careful measurements on how much the path of the cathode ray was bent in a magnetic field and how much energy they carried. From this work Thomson could describe the mass/charge ratio of the cathode ray particles. *What amazing result did Thomson find?*

**Thomson's Atomic Model:** Thomson presented three hypotheses from his experiments. Only two were accepted by physicists – in fact the third was shown to be wrong! From the first two came a model of the atom known as the *Plum Pudding* model. Complete the atom drawing below to illustrate Thomson's plum pudding model. Explain how this fits with his observations.



# Modeling Chemistry Amdt 2013 U6 Ws1 V3

**Peipei Pang**



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## **Table of Contents Modleing Chemistry Amta 2013 U6 Ws1 V3**

1. Understanding the eBook Modleing Chemistry Amta 2013 U6 Ws1 V3
  - The Rise of Digital Reading Modleing Chemistry Amta 2013 U6 Ws1 V3
  - Advantages of eBooks Over Traditional Books
2. Identifying Modleing Chemistry Amta 2013 U6 Ws1 V3
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Modleing Chemistry Amta 2013 U6 Ws1 V3
  - User-Friendly Interface
4. Exploring eBook Recommendations from Modleing Chemistry Amta 2013 U6 Ws1 V3
  - Personalized Recommendations
  - Modleing Chemistry Amta 2013 U6 Ws1 V3 User Reviews and Ratings
  - Modleing Chemistry Amta 2013 U6 Ws1 V3 and Bestseller Lists
5. Accessing Modleing Chemistry Amta 2013 U6 Ws1 V3 Free and Paid eBooks
  - Modleing Chemistry Amta 2013 U6 Ws1 V3 Public Domain eBooks
  - Modleing Chemistry Amta 2013 U6 Ws1 V3 eBook Subscription Services
  - Modleing Chemistry Amta 2013 U6 Ws1 V3 Budget-Friendly Options

6. Navigating Modleing Chemistry Amta 2013 U6 Ws1 V3 eBook Formats
  - ePub, PDF, MOBI, and More
  - Modleing Chemistry Amta 2013 U6 Ws1 V3 Compatibility with Devices
  - Modleing Chemistry Amta 2013 U6 Ws1 V3 Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Modleing Chemistry Amta 2013 U6 Ws1 V3
  - Highlighting and Note-Taking Modleing Chemistry Amta 2013 U6 Ws1 V3
  - Interactive Elements Modleing Chemistry Amta 2013 U6 Ws1 V3
8. Staying Engaged with Modleing Chemistry Amta 2013 U6 Ws1 V3
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Modleing Chemistry Amta 2013 U6 Ws1 V3
9. Balancing eBooks and Physical Books Modleing Chemistry Amta 2013 U6 Ws1 V3
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Modleing Chemistry Amta 2013 U6 Ws1 V3
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Modleing Chemistry Amta 2013 U6 Ws1 V3
  - Setting Reading Goals Modleing Chemistry Amta 2013 U6 Ws1 V3
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Modleing Chemistry Amta 2013 U6 Ws1 V3
  - Fact-Checking eBook Content of Modleing Chemistry Amta 2013 U6 Ws1 V3
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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