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CHAPTER 6 REVIEW

# The Periodic Law

# SECTION 3

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	_ When an electron is added to		
	(a) always absorbed.	<ul><li>(c) either absorbed or release</li></ul>	
	(b) always released.	<ul><li>(d) seither also rised nor rele</li></ul>	annech.
-	atom's	ove an electron from a neutral atom i	is the
	(a) electron affinity.	<ul><li>(c) electronegativity.</li></ul>	
	(b) electron energy.	(d) neither absorbed nor rele	an act.
A. Proper	a left to right across a period o	n the periodic sable,	
		<ul> <li>affinity values tend to become more ve or positive).</li> </ul>	
	b. to nazioni checonomi	iors emergy values tend to (Increase o e).	ř.
	e, atomie	radii teral to become (larger or trnall	See 13
٠	_a.Name the baloges with the	e least-negative electron adfinity.	
	b. Name the alkali metal wit	h the highest ionization energy.	
	_ c.Name the element in Perio	ed 3 with the smallest atomic radius.	
	d Name the Group 14 eleme	nt with the longest electronegativity.	
a. Webste	e the electron configuration of	the following:	
a. No			
b. No	3.		
c. O			
d. O			
6.0	21.0		

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**Peipei Pang** 

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