

- Chapter 5 Assessment Problems pp. 146-147 #33, 34, 39, 44, 45, 47, 49, 50, 52, 53, 56, 57, 59, 61, 63, 64, 65-72, 76, 78, 79, 80, 81(a-d,f), 82.
- 33. Explain how the gaseous neon atoms in a neon sign emit light.
Electricity passing through the tube excites electrons in the neon atoms to higher energy levels. When the excited electrons relax, they emit light as they lose energy.
- 34. What is a photon?
A particle of electromagnetic radiation having a mass of zero and carrying a quantum of energy.
- 39. What is the difference between an atom's ground state and an excited state?
The ground state is the lowest energy state, while any energy state higher than that is excited.
- 44. What is an atomic orbital?
A three-dimensional region around the nucleus describing an electron's probably location.
- 45. What is the probability that an electron will be found within an atomic orbital? 90%
- 47. How many energy sublevels are contained in each of the hydrogen atoms first three energy levels?
Level 1: 1; level 2: 2; level 3: 3
- 49. Which of the following atomic orbital designations are incorrect?
a. 7f b. 3f c. 2d d. 6p b & c are incorrect—the orbitals do not exist
- 50. What do the designations s, p, d, and f specify with respect to the atom's orbitals?
The shapes of the orbitals.
- 52. What is the maximum number of electrons an orbital may contain? 2
- 53. Why is it impossible to know precisely the velocity and position of an electron at the same time?
The photon required to measure an electron's velocity or position changes both the position and velocity of the electron (like a bowling ball thrown at a ping pong ball).
- 56. How is an orbital's principal quantum number related to the atom's major energy levels?
They are the same.
- 57. Explain the meaning of the Aufbau principle as it applies to atoms with many electrons.
The Aufbau principle describes the sequence in which an atom's orbitals fill with electrons.
- 59. Why must the two arrows within a single block of an orbital diagram be written in opposite (up and down) directions?
They must have opposite spins.
- 61. What are valence electrons? How many of a magnesium atom's 12 electrons are valence electrons?
The electrons in an atom's outermost orbitals. Magnesium has 2 electrons.
- 63. Describe the difference between a quantum and a photon.
A quantum is the minimum amount of energy that can be lost or gained by an atom, while a photon is a particle of light that carries a quantum of energy.

Modern Chemistry Chapter 5 Review Answer Key

WJ Hussar



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Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual Leo J. Malone, Theodore O.

Dolter, 2012-01-03 The 9th edition of Malone s Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections assessment exercises at the end each section and relevant chapter problems at the end of each chapter A new Math Check allows quick access to the needed basic skill The first chapter now includes brief introductions to several fundamental chemical concepts and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem Every concept in the text is clearly illustrated with one or more step by step examples Making it Real essays have been updated to present timely and engaging real world applications emphasizing the relevance of the material they are learning This edition continues the end of chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter

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molecules which highlights common problems made by scientists **Merrill Chemistry** Robert C. Smoot, Smoot, Richard G. Smith, Jack Price, 1998 *Saturday Review*, 1863 Introduction to General, Organic, and Biochemistry Morris Hein, Scott Pattison, Susan Arena, Leo R. Best, 2014-01-15 The most comprehensive book available on the subject Introduction to General Organic and Biochemistry 11th Edition continues its tradition of fostering the development of problem solving skills featuring numerous examples and coverage of current applications Skillfully anticipating areas of difficulty and pacing the material accordingly this readable work provides clear and logical explanations of chemical concepts as well as the right mix of general chemistry organic chemistry and biochemistry An emphasis on real world topics lets readers clearly see how the chemistry will apply to their career Basic Concepts of Chemistry, Study Guide Leo J. Malone, 2000 The text s three main goals are to introduce chemistry as a living relevant science to encourage learning and critical thinking and to help readers overcome the math difficulties that impede their progress in chemistry Designed to help readers master the principles of general chemistry As a prep book it promotes active involvement with the material There are special features throughout that reinforce concepts and help to develop strong problem solving and study skills Updated to Include an Interactive Learning Ware problems CD containing several of the chapter ending problems from the book in an interactive tutorial with feedback to help readers set up and solve problems **Foundations of College Chemistry, Alternate** Morris Hein, Susan Arena, 2010-01-26 Learning the fundamentals of chemistry can be a difficult task to undertake for health professionals For over 35 years this book has helped them master the chemistry skills they need to succeed It provides them with clear and logical explanations of chemical concepts and problem solving They ll learn how to apply concepts with the help of worked out examples In addition Chemistry in Action features and conceptual questions checks brings together the understanding of chemistry and relates chemistry to things health professionals experience on a regular basis *Introduction to Modern Chemistry: Student Manual* Eugene Meyer, 1979 **Operation of Wastewater Treatment Plants**, 2004 **OAT 2017-2018 Strategies, Practice & Review with 2 Practice Tests** Kaplan Test Prep, 2016-10-04 Kaplan s OAT 2017 2018 Strategies Practice Review provides the content review test taking strategies and realistic practice you need to get the OAT results you want Updated for the latest test changes OAT 2017 2018 is your guide to facing Test Day with confidence The Best Review Two full length online practice tests More than 600 practice questions for every subject with detailed answers and explanations 16 page tear out full color study sheets for quick review on the go A guide to the current OAT Blueprint so you know exactly what to expect on Test Day Comprehensive review of all of the content covered on the OAT Biology General Chemistry Organic Chemistry Reading Comprehension Physics Quantitative Reasoning Kaplan s proven strategies for Test Day success Expert Guidance Kaplan s expert psychometricians ensure our practice questions and study materials are true to the test We invented test prep Kaplan www.kaptest.com has been helping students for almost 80 years Our proven strategies have helped legions of students achieve their dreams *The Saturday Review of Politics, Literature, Science and Art*, 1860

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