

Chapter 15

Oxidation-Reduction Reactions

SECTION 15.1

Transfer electrons answer the following questions for the species presented.

1. _____ Add the following oxidation numbers and show whether it's correct.
a. $\text{Cu}^{+2} + \text{H}^{+1} + \text{O}^{-2} + \text{H}^{+1} = 0$
b. $\text{Fe}^{+3} + \text{O}^{-2} + \text{H}^{+1} + \text{O}^{-2} = 0$
c. $\text{Mg}^{+2} + \text{O}^{-2} + \text{H}^{+1} = 0$

2. Assign oxidation numbers to each element in the compounds listed in the table below.

- | | |
|-------|--|
| _____ | a. H_2O for H_2O |
| _____ | b. H_2SO_4 for H_2SO_4 |
| _____ | c. CO_2 for CO_2 |
| _____ | d. H_2SO_3 for H_2SO_3 |
| _____ | e. H_2SO_4 for H_2SO_4 |
| _____ | f. H_2SO_4 for H_2SO_4 |
| _____ | g. H_2SO_4 for H_2SO_4 |

3. In each of the following half-reactions, determine the oxidation state.

- | | |
|-------|---|
| _____ | a. $\text{Fe}^{+2} + \text{O}^{-2} = 0$ |
| _____ | b. $\text{Fe}^{+2} + \text{O}^{-2} = 0$ |
| _____ | c. $\text{Fe}^{+2} + \text{O}^{-2} = 0$ |

4. _____ a. $\text{Fe}^{+2} + \text{O}^{-2} = 0$ for $\text{Fe}^{+2} + \text{O}^{-2} = 0$

5. _____ b. $\text{Fe}^{+2} + \text{O}^{-2} = 0$ for $\text{Fe}^{+2} + \text{O}^{-2} = 0$

6. _____ c. $\text{Fe}^{+2} + \text{O}^{-2} = 0$ for $\text{Fe}^{+2} + \text{O}^{-2} = 0$

- | | |
|-------|--|
| _____ | a. $\text{Fe}^{+2} + \text{O}^{-2} = 0$ for $\text{Fe}^{+2} + \text{O}^{-2} = 0$ |
|-------|--|

7. _____ b. $\text{Fe}^{+2} + \text{O}^{-2} = 0$ for $\text{Fe}^{+2} + \text{O}^{-2} = 0$

8. _____ c. $\text{Fe}^{+2} + \text{O}^{-2} = 0$ for $\text{Fe}^{+2} + \text{O}^{-2} = 0$

9. _____ d. $\text{Fe}^{+2} + \text{O}^{-2} = 0$ for $\text{Fe}^{+2} + \text{O}^{-2} = 0$

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