

## Reviewing Content

42. The solvent is the substance in which the solute is dissolved.
43. Random collisions of the solvent molecules with the solute particles provide enough force to overcome gravity.
44. **solubility**: the amount of a substance that dissolves in a given quantity of solvent at specified conditions of temperature and pressure to produce a saturated solution.  
**saturated solution**: a solution containing the maximum amount of solute for a given amount of solvent at a constant temperature and pressure. **unsaturated solution**: a solution that contains less solute than a saturated solution at a given temperature and pressure. **miscible**: describes liquids that dissolve in each other. **immiscible**: describes liquids that are insoluble in each other.
45. Particles of solute crystallize.
46. No; if there were undissolved solute, the excess solute would come out of a supersaturated solution.
47.  $5.55 \times 10^2$  g  $\text{AgNO}_3$
48. Solubility increases with pressure.
49. a.  $1.6 \times 10^{-2}$  g/L  
b.  $4.7 \times 10^{-2}$  g/L
50. *Dilute* and *concentrated* are relative terms and are not quantitative. Molarity provides the exact number of moles of solute per liter of solution.
51. Molarity is the number of moles of solute dissolved in one liter of solution.  
a. 1.3M KCl  
b.  $3.3 \times 10^{-1}$ M  $\text{MgCl}_2$
52.  $2.00 \times 10^3$  mL
53. a.  $5.0 \times 10^{-3}$  mol NaCl, 29 g NaCl  
b. 1.0 mol  $\text{KNO}_3$ ,  $1.0 \times 10^{-2}$  g  $\text{KNO}_3$   
c.  $2.5 \times 10^3$  mol  $\text{CaCl}_2$ , 2.8 g  $\text{CaCl}_2$
54. a.  $2.3 \times 10^1$  g NaCl  
b. 2.0 g  $\text{MgCl}_2$
55. a. 16% (v/v) ethanol  
b. 63.6% (v/v) isopropyl alcohol
56. Colligative properties are properties of a solution that depend only on the number of solute particles; boiling-point elevation, freezing-point depression, and vapor-pressure lowering. Boiling points are elevated because shells of solvent form around solute particles, reducing the amount of solvent molecules that have sufficient energy to escape the solution; relative to the pure solvent, the amount of energy required to cause vaporization or boiling increases. Solutes disrupt the ordering of the solvent structure, so more kinetic energy must be withdrawn from a solution for it to solidify. This lowers the freezing point of the solution.
57. a. sea water  
b. 1.5M  $\text{KNO}_3$   
c. 0.100M  $\text{MgCl}_2$
58. The effective molality of the  $\text{Ca}(\text{NO}_3)_2$  solution is 3m. The effective molality of the  $\text{NaNO}_3$  solution is 2m.
59. When vapor pressure is lowered relative to pure solvent, more energy must be supplied to reach the boiling point; thus the boiling point is increased relative to pure solvent.
60. The salt lowers the freezing point of the ice-water cooling mixture.
61. 1M solution: 1 mol of solute in 1 L of solution; 1m solution: 1 mol of solute in 1000 g of solvent
62. Add 27.0 g  $\text{H}_2\text{O}$  to 32.0 g  $\text{CH}_3\text{OH}$ .
63. a.  $100.26^\circ\text{C}$   
b.  $101.54^\circ\text{C}$
64. a.  $-4.46^\circ\text{C}$   
b.  $-2.2^\circ\text{C}$
65. a.  $-1.1^\circ\text{C}$   
b.  $-0.74^\circ\text{C}$   
c.  $-1.5^\circ\text{C}$

## Understanding Concepts

66. a. The freezing-point depression is twice as great for solute B; solute B must provide twice as many particles in solution.

# Modern Chemistry Chapter 16 Answers

**C. A. Trapp, Peter Atkins, Julio dePaula**



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**Principles of Modern Chemistry** David W. Oxtoby, H. Pat Gillis, Laurie J. Butler, 2016-01-01 Long considered the standard for honors and high level mainstream general chemistry courses PRINCIPLES OF MODERN CHEMISTRY continues to set the standard as the most modern rigorous and chemically and mathematically accurate text on the market This authoritative text features an atoms first approach and thoroughly revised chapters on Quantum Mechanics and Molecular Structure Chapter 6 Electrochemistry Chapter 17 and Molecular Spectroscopy and Photochemistry Chapter 20 In addition the text utilizes mathematically accurate and artistic atomic and molecular orbital art and is student friendly without compromising its rigor End of chapter study aids focus on only the most important key objectives equations and concepts making it easier for students to locate chapter content while applications to a wide range of disciplines such as biology chemical engineering biochemistry and medicine deepen students understanding of the relevance of chemistry beyond the classroom

*Modern Chemistry* Nicholas D. Tzimopoulos, 1993 **Chemistry in Your Life Solutions Manual** Joseph Laurino, 2005-12-16 Designed to help students understand the material better and avoid common mistakes Includes solutions and explanations to odd numbered exercises

**Modern Chemistry** Holt Rinehart & Winston, Holt, Rinehart and Winston Staff, 2001

The Development of Modern Chemistry Aaron J. Ihde, 1984-01-01 From ancient Greek theory to the explosive discoveries of the 20th century this authoritative history shows how major chemists their discoveries and political economic and social developments transformed chemistry into a modern science 209 illustrations 14 tables Bibliographies Indices

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**Fragrance Chemistry** Charles S. Sell, 2019-03-26 Comprehensively teaches all of the fundamentals of fragrance chemistry Ernest Beaux the perfumer who created Chanel No 5 said One has to rely on chemists to find new aroma chemicals creating new original notes In perfumery the future lies primarily in the hands of chemists This book provides chemists and chemists to be with everything they need to know in order to create welcome new fragrances for the world to enjoy It offers a simplified introduction into organic chemistry including separation techniques and analytical methodologies discusses the structure of perfume creation with respect to the many reactive ingredients in consumer products and shows how to formulate effective and long lasting scents Fundamentals of Fragrance Chemistry starts by covering the structure of matter in order to show how its building blocks are held together It continues with chapters that look at hydrocarbons and heteroatoms A description of the three states of matter and how each can be converted into another is offered next followed by coverage of separation and purification of materials Other chapters examine acid base reactions oxidation and reduction reactions perfume structure the mechanism of olfaction natural and synthetic fragrance ingredients and much more Concentrates on aspects of organic chemistry which are of particular importance to the fragrance industry Offers non chemists a simplified yet complete introduction to organic chemistry from separation techniques and analytical

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