## Chapter 4 &5: Modern Model of the Atom and Bonding

Leader: Katie
Course: Chem 163
Instructor: Burnett/Kingston
Date: September 28, 2009

Supplemental Instruction lowa State University

Review

- 1. What is the noble gas configuration of the following atoms and ions?
  - a) Ar [Ne] 352 3pf or [Ar]
  - b) Nan+DAO 25 20 5 Blo)
  - e) C^2-[He] 26-204
  - d) 0"2-[He] 25" 2p" [NE]
  - e) Car2+ Lue > 35 50 60 CAr3

New Material

- 1. Define the following terms and answer associated questions.
  - a) Heisenberg uncertainty principle: Principle that states that the position and momentum of a subatanic particle connot be known simultaneously.
  - b) orbital: 310 volume in space, cloud like distribution of electrons.

-What is its maximum capacity? ⊇ €-

c) covalent bond: Chemical band where the electrons are shared.

-What does it represent? The covalent-band represents on evergy minimum,
-Is energy absorbed or released when a covalent bond is formed? Evergy is released.

d) bond length:

-What causes bond length? +ne attraction of electrons to the nuclei controps and fanctei;
-What units is bond length measured in? A = Angreen = 10 = X = 1m

2. How many orbitals are there in any s sublevel? Any p sublevel? Any d sublevel?

S=X Ichital - 2cp = 3orbitals - 6ea = Sorbitals - 10e-

3. Draw the orbitals for the 2p^5. How many electrons are there in each orbital?

8 35 2



- 4. Write the following atoms in Lewis Dot notation.
  - a) Nb) H
- 63 NJ-

O Xe

- c) Xe d) Fe
- @ H.

⊕ Fe

# **Modern Chemistry Chapter 12 3 Review Answers**

Zhe Xu

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