

## Chapter 2

1. A compound is a combination of 2 or more different elements.
2. Oxygen, Carbon, Hydrogen, Nitrogen
3. **Trace Element**- Required by an organism only in small quantities, such as iodine.
4. They are not accurate representations because the nucleus is not to scale, it is actually much smaller. And the electrons are orbiting the nucleus in a cloud of negative charge.
5. a. Neutron and Proton are both part of the nucleus, proton is positive charge, neutron is neutral.  
b. Atomic number is the number of protons, while mass number is the sum of protons plus neutrons in the atom.  
c. Atomic mass is the weight of both the protons and neutrons, electron weight is so small that it is negligible.
6. You can subtract the atomic number from the mass number to find the number of protons.
7. **Isotopes**- Same element, so the same number of protons, but they have a different number of neutrons. There can be stable or radioactive isotopes (radioactive has the nucleus decay spontaneously).
8. Used to trace biological activities and also to date fossils.
9. **Energy**- The capacity to cause change.  
**Potential Energy**- The energy that matter has because of its location or structure. An electron's state of potential energy is called its energy level, or **electron shell**.  
**Valence Electrons**- Outermost shell. Elements with a full valence shell are chemically inert.  
**Electrons at a higher shell have more potential energy.**
10. **Nonpolar Covalent Bond**- The atoms share the electron equally.  
**Polar Covalent Bond**- One atom is more electronegative, and the atoms do not share the electrons equally.  
**Ionic Bonds**- A cation and anion are oppositely charged, so they attract each other. ( $\text{Na}^+\text{Cl}^-$ )

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