4. It is not safe to put seroed canisters in a campfire, because the pressure inside the canisters gets very high and they can explode. If I have a 1.0 liter canister that holds 2 moles of gas, and the campfire temperature is 1400° C, what is the pressure inside the canister?
What equation should be used?
Solver
5. If I place 3 moles of N _s and 4 moles of O _s in a 35 L container at a temperature of 25° C, what will the pressure of the resulting modure of pases he? What equation(s) should be used?
Solve:
6. At a pressure of 1, 0 atm and a temp of 298K, calcium carbonate decomposes to form carbon dioxide and calcium carde. How many grams of calcium carbonate will be needed to form 3.45 liters of carbon dioxide?
Review (stid) - student will apply Kinetic Holecular Theory when describing a novel, real- world elituation in terms of collisions and motions of particles). 7. A married empty bottle has a balloon attached to its top, as done in station 1 of Expt #0.05206. When the bottle is placed into a cold water bath, the balloon deflatests pulled in to the bottle. Using the finetic Holecular theory, explain why this occurs, making sure to refer to the motion of the gas malecules.

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