

CHAPTER 6 REVIEW

The Periodic Law

SECTION 3

SHORT ANSWER Answer the following questions in the space provided.

1. _____ When an electron is added to a neutral atom, energy is
 - (a) always absorbed.
 - (b) always released.
 - (c) either absorbed or released.
 - (d) neither absorbed nor released.

2. _____ The energy required to remove an electron from a neutral atom is the atom's
 - (a) electron affinity.
 - (b) electron energy.
 - (c) electronegativity.
 - (d) neither absorbed nor released.

3. From left to right across a period on the periodic table,
 - _____ a. electron affinity values tend to become more (negative or positive).
 - _____ b. ionization energy values tend to (increase or decrease).
 - _____ c. atomic radii tend to become (larger or smaller).

4. _____ a. Name the halogen with the least-negative electron affinity.
 _____ b. Name the alkali metal with the highest ionization energy.
 _____ c. Name the element in Period 3 with the smallest atomic radius.
 _____ d. Name the Group 14 element with the longest electronegativity.

5. Write the electron configuration of the following:
 - a. Na

 - b. Na⁺

 - c. Cl

 - d. Cl⁻

 - e. Cu²⁺

Modern Chemistry Review Answers Section

L Reisser



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Table of Contents Modern Chemistry Review Answers Section

1. Understanding the eBook Modern Chemistry Review Answers Section
 - The Rise of Digital Reading Modern Chemistry Review Answers Section
 - Advantages of eBooks Over Traditional Books
2. Identifying Modern Chemistry Review Answers Section
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Modern Chemistry Review Answers Section
 - User-Friendly Interface
4. Exploring eBook Recommendations from Modern Chemistry Review Answers Section
 - Personalized Recommendations
 - Modern Chemistry Review Answers Section User Reviews and Ratings
 - Modern Chemistry Review Answers Section and Bestseller Lists

5. Accessing Modern Chemistry Review Answers Section Free and Paid eBooks
 - Modern Chemistry Review Answers Section Public Domain eBooks
 - Modern Chemistry Review Answers Section eBook Subscription Services
 - Modern Chemistry Review Answers Section Budget-Friendly Options
6. Navigating Modern Chemistry Review Answers Section eBook Formats
 - ePub, PDF, MOBI, and More
 - Modern Chemistry Review Answers Section Compatibility with Devices
 - Modern Chemistry Review Answers Section Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Modern Chemistry Review Answers Section
 - Highlighting and Note-Taking Modern Chemistry Review Answers Section
 - Interactive Elements Modern Chemistry Review Answers Section
8. Staying Engaged with Modern Chemistry Review Answers Section
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Modern Chemistry Review Answers Section
9. Balancing eBooks and Physical Books Modern Chemistry Review Answers Section
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Modern Chemistry Review Answers Section
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Modern Chemistry Review Answers Section
 - Setting Reading Goals Modern Chemistry Review Answers Section
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Modern Chemistry Review Answers Section
 - Fact-Checking eBook Content of Modern Chemistry Review Answers Section
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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