CHAPTER 8 REVIEW

Chemical Equations and Reactions

SECTION 8-1

SHORT ANSWER Answer the following questions in the space provided.

1. Match the symbol on the left with its appropriate description on the right.

_____ A

(a) A precipitate forms.

(b) A gas forms.

_____ ↑

(c) A reversible reaction occurs.

....(1)

(d) Heat is applied to the reactants.

____ (aq)

(e) A chemical is dissolved in water.
(f) A chemical is in the liquid state.

2. Finish balancing the following equation:

 $3\text{Fe}_3\text{O}_4 + \underline{\hspace{1cm}} \text{Al} \rightarrow \underline{\hspace{1cm}} \text{Al}_2\text{O}_3 + \underline{\hspace{1cm}} \text{Fe}$

3. In each of the following formulas with coefficients, write the total number of atoms present.

_____ a. 4SO₂

_____ b. SO₂

_____ c. 3Al₂(SO₄)₃

_____ d, 6×10^{23} HNO₃

- Convert the following word equation into a balanced chemical equation: aluminum metal + copper(II) fluoride → aluminum fluoride + copper metal
- 5. One way to test the salinity of a water supply is to add a few drops of silver nitrate of known concentration to the water. As the solutions of sodium chloride and silver nitrate mix, a precipitate of silver chloride forms, leaving sodium nitrate in solution. Translate these sentences into a balanced chemical equation.
- a. Balance the following equation: NaHCO_y(s)
 [∆]⇒ Na₂CO_y(s) + H₂O(g) + CO₂(g)

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Modern Chemistry Chapter 8 Review Answers

Lionello Pogliani,Francisco Torrens,A. K. Haghi

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