

Chapter 6 Practice Test: Chemical Bonding

In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question.

- _____ 1. The charge on an ion is
a. always positive.
b. always negative.
c. either positive or negative.
d. zero.
- _____ 2. According to the octet rule, a calcium atom has a tendency to
a. lose one electron.
b. lose two electrons.
c. gain one electron.
d. gain two electrons.
- _____ 3. If a compound forms by ionic bonding, which is *not* true?
a. A positively charged atom or group of atoms attracts a negatively charged atom or group of atoms.
b. The net charge of the compound is zero.
c. The compound contains just two atoms, each of opposite charge.
d. Several ions group together in a tightly packed structure.
- _____ 4. The only property listed that is *not* characteristic of ionic compounds is
a. high melting point.
b. hardness.
c. lack of crystal structure.
d. good conductor of electricity.
- _____ 5. Which formula listed below represents a polyatomic ion?
a. NH_4^+
b. H_2SO_4
c. NaCl
d. H_2O
- _____ 6. The melting points of ionic compounds are higher than the melting points of molecular compounds because
a. ionic substances tend to vaporize at room temperature.
b. ionic substances are brittle.
c. attractive forces between ions are greater than the attractive forces between molecules.
d. the numbers of positive and negative charges are equal in an ionic compound.

Modern Chemistry Answers Review

Liyong Dong



Modern Chemistry Answers Review:

Modern Chemistry Holt Rinehart & Winston, Holt, Rinehart and Winston Staff, 2001 **Reviews of Modern Quantum Chemistry** Kali Das Sen, 2002 This important book collects together state-of-the-art reviews of diverse topics covering almost all the major areas of modern quantum chemistry. The current focus in the discipline of chemistry synthesis, structure, reactivity, and dynamics is mainly on control. A variety of essential computational tools at the disposal of chemists have emerged from recent studies in quantum chemistry. The acceptance and application of these tools in the interfacial disciplines of the life and physical sciences continue to grow. The new era of modern quantum chemistry throws up promising potentialities for further research. *Reviews of Modern Quantum Chemistry* is a joint endeavor in which renowned scientists from leading universities and research laboratories spanning 22 countries present 59 in-depth reviews. Along with a personal introduction written by Professor Walter Kohn, Nobel laureate in Chemistry, 1998, the articles celebrate the scientific contributions of Professor Robert G. Parr on the occasion of his 80th birthday. List of Contributors: W. Kohn, M. Levy, R. Pariser, B. R. Judd, E. Lo, B. N. Plakhutin, A. Savin, P. Politzer, P. Lane, J. S. Murray, A. J. Thakkar, S. R. Gadre, R. F. Nalewajski, K. Jug, M. Randić, G. Del Re, U. Kaldor, E. Eliav, A. Landau, M. Ehara, M. Ishida, K. Toyota, H. Nakatsuji, G. Maroulis, A. M. Mebel, S. Mahapatra, R. Carbó, C. Dorca, U. Nagy, I. A. Howard, N. H. March, S. O. Collier, B. Liu, R. G. Pearson, N. Watanabe, S. Ten-no, S. Iwata, Y. Udagawa, E. Valderrama, X. Fradera, I. Silanes, J. M. Ugalde, R. J. Boyd, E. V. Ludea, V. V. Karasiev, L. Massa, T. Tsuneda, K. Hirao, J. M. Tao, J. P. Perdew, O. V. Gritsenko, M. Gruning, E. J. Baerends, F. Aparicio, J. Garza, A. Cedillo, M. Galvin, R. Vargas, E. Engel, A. Hack, R. N. Schmid, R. M. Dreizler, J. Poater, M. Sola, M. Duran, J. Robles, X. Fradera, P. K. Chattaraj, A. Poddar, B. Maiti, A. Cedillo, S. Gutierrez, O. Coliva, P. Jaque, A. Toro-Labbé, R. H. Chermette, P. Boulet, S. Portmann, P. Fuentealba, R. Contreras, P. Geerlings, F. De Proft, R. Balawender, D. P. Chong, A. Vela, G. Merino, F. Kootstra, P. L. de Boeij, R. van Leeuwen, J. G. Snijders, N. T. Maitra, K. Burke, H. Appel, E. K. U. Gross, M. K. Harbola, H. F. Hameka, C. A. Daul, I. Ciofini, A. Bencini, S. K. Ghosh, A. Tachibana, J. M. Cabrera, O. Trujillo, F. Tenorio, O. Mayorga, M. Cases, V. Kumar, Y. Kawazoe, A. M. Koster, P. Calaminici, Z. Gmez, U. Reveles, J. A. Alonso, L. M. Molina, M. J. Lopez, F. Dugue, A. Maanes, C. A. Fahlstrom, J. A. Nichols, D. A. Dixon, P. A. Derosa, A. G. Zacarias, J. M. Seminario, D. G. Kanhere, A. Vichare, S. A. Blundell, Z. CoY, Lu, H. CoY, Liu, M. Elstner, W. CoT, Yang, J. Muoz, X. Fradera, M. Orozco, F. J. Luque, P. Tarakeshwar, H. M. Lee, K. S. Kim, M. Valiev, E. J. Bylaska, A. Gramada, J. H. Weare, J. Brickmann, M. Keil, T. E. Exner, M. Hoffmann. Probability Distributions and Valence Shells in Atoms, A. Savin. Information Theoretical Approaches to Quantum Chemistry, S. R. Gadre. Quantum Chemical Justification for Clar's Valence Structures, M. Randić. Functional Expansion Approach in Density Functional Theory, S. B. Liu. Norm-conserving Pseudopotentials for the Exact Exchange Functional, E. Engel et al. Volume II: Chemical Reactivity and Dynamics within a Density-based Quantum Mechanical Framework, P. K. Chattaraj et al. Fukui Functions and Local Softness, H. Chermette et al. The Nuclear Fukui Function, P. Geerlings et al. Causality in Time-Dependent Density Functional Theory, M. K. Harbola. Theoretical Studies of Molecular Magnetism, H. F. Hameka. Melting in Finite Sized Systems, D. G. Kanhere et al. Density

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