Atomic Structure – Study Guide Grade S – Physical Science

Atom -- smallest particle of an element that has all the properties of that element.

3 main subatomic particles:

- I. protons
- 2. neutrons
- 3. electrons



nucleus -- core of the atom, it is positively charged 99.9 % of the atoms mass is in the nucleus.

Protons and neutrons are found in the nucleus.

Protons — have a positive (+) charge. Found in the nucleus of the atom.

The number of protons determines what element an atom is.

Example: Carbon has 6 protons

Atomic Number - the number of protons, each element has it own unique number of protons.

 Neutrons — a neutral charge (no charge) found in the nucleus. These determine whether an element is radioactive.

Protons and neutrons are about ovar atomic mass unit (annu). Electrons have a much smaller mass — it takes almost 2000 electrons to equal 1 annu.

Atomic Mass - the total number of protons and neutrons.

Mass Number - Atomic Mass that is rounded.

To find just how many neutrons an atom has:

neutrons = atomic mass - atomic

Isotopes - atoms of the same element that have different number of neutrons.

Carbon-12 Carbon-14
6 protons 6 protons
6 neutrons 8 neutrons

NOTICE - the number of protons DOES NOT change if it is to be carbon -- only the number of neutrons.

3. Electrons - have a negative (-) charge. Found moving around the atom is a space called orbital's (shells). Those electrons farthest away from the nucleus are called <u>valence electrons</u>, and are involved in the formation of chemical bonds.

In an uncharged atom the number of protons is equal to the number of electrons.

lon - when an atom gains or loses an electron (has a charge)

If an atom gains electrons = negative (-) charge

If an atom loses electrons - positive (+) charge

How to find:

protons - atomic

neutrons = atomic mass - atomic

electrons = atomic # — charge (if any)

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